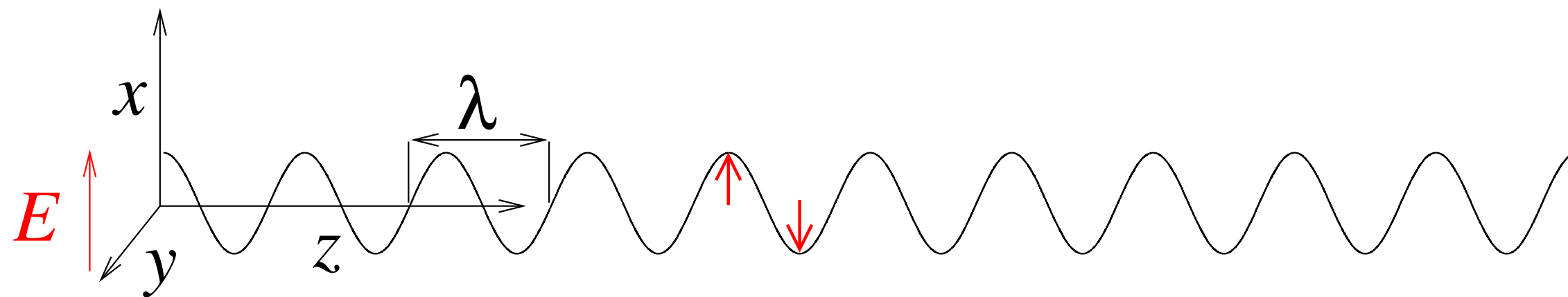
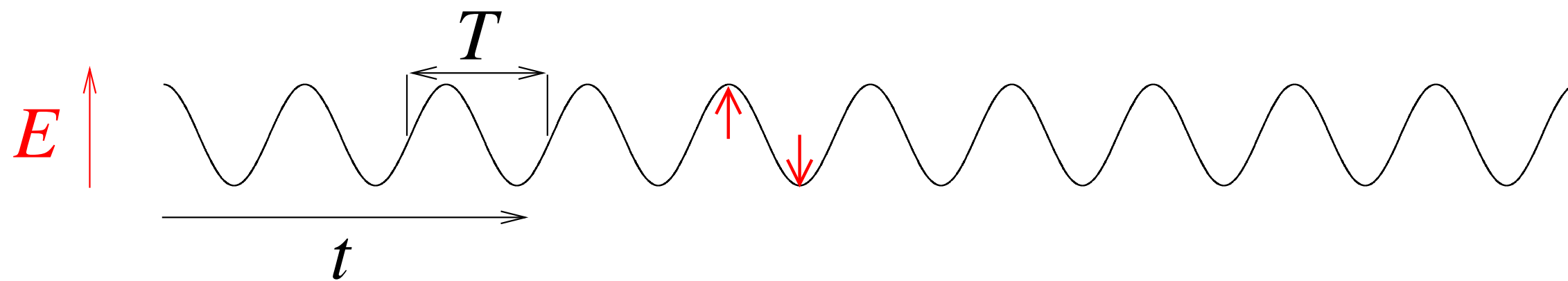
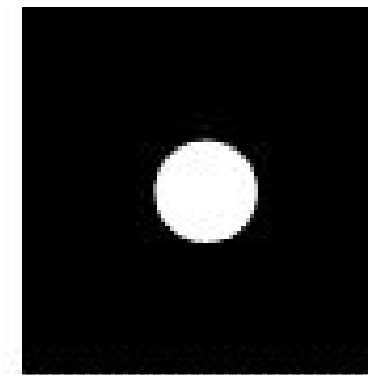
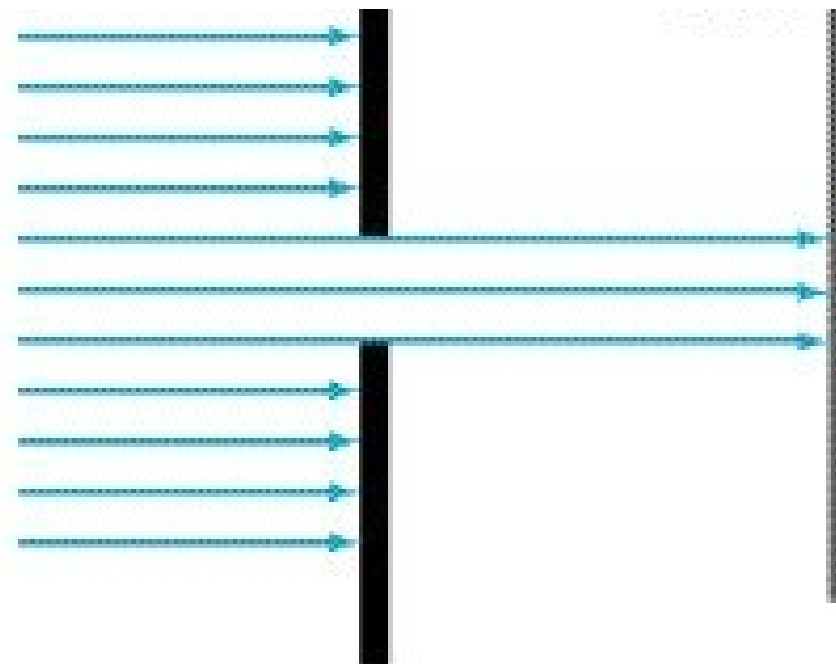


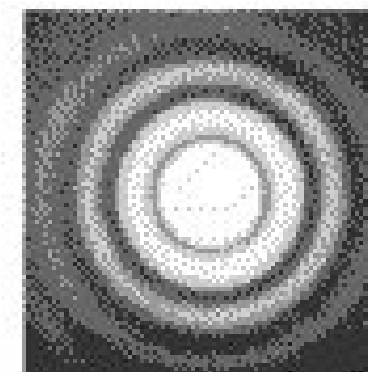
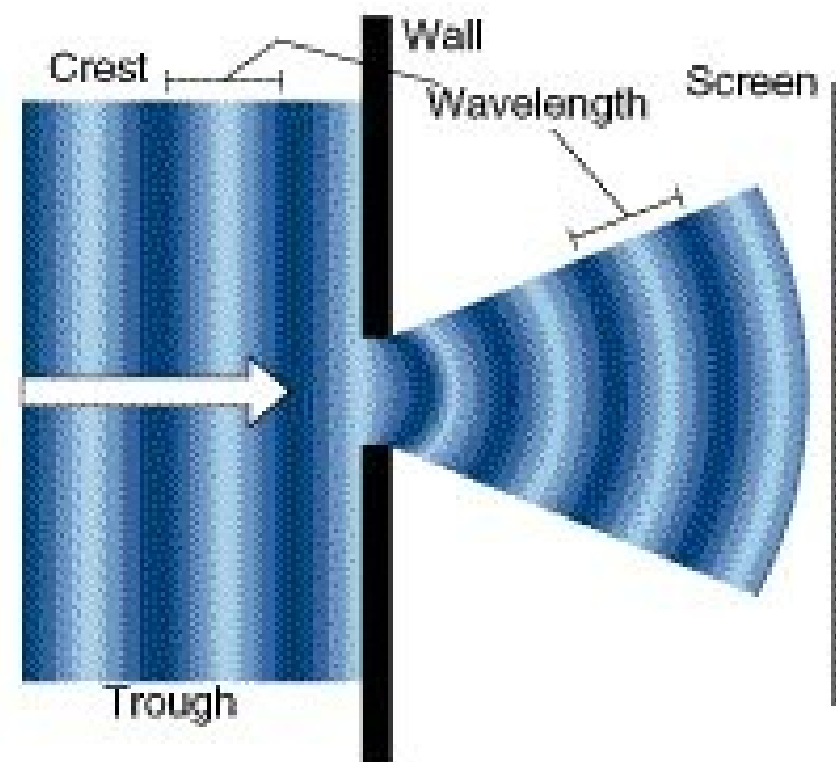
OPTICAL METHODS



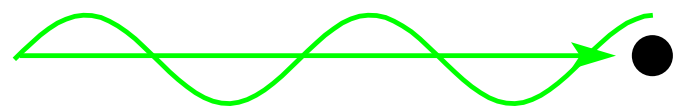


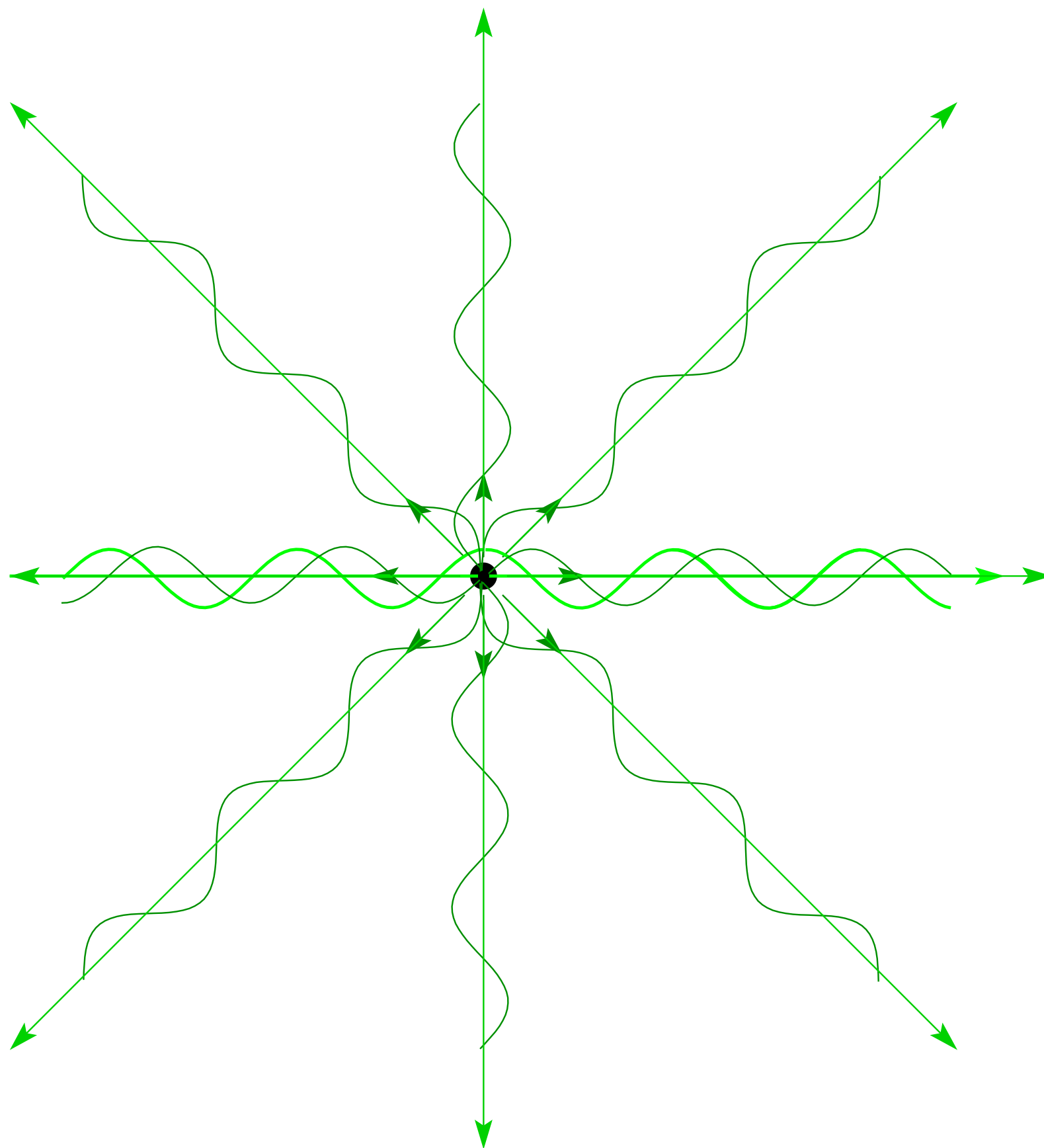


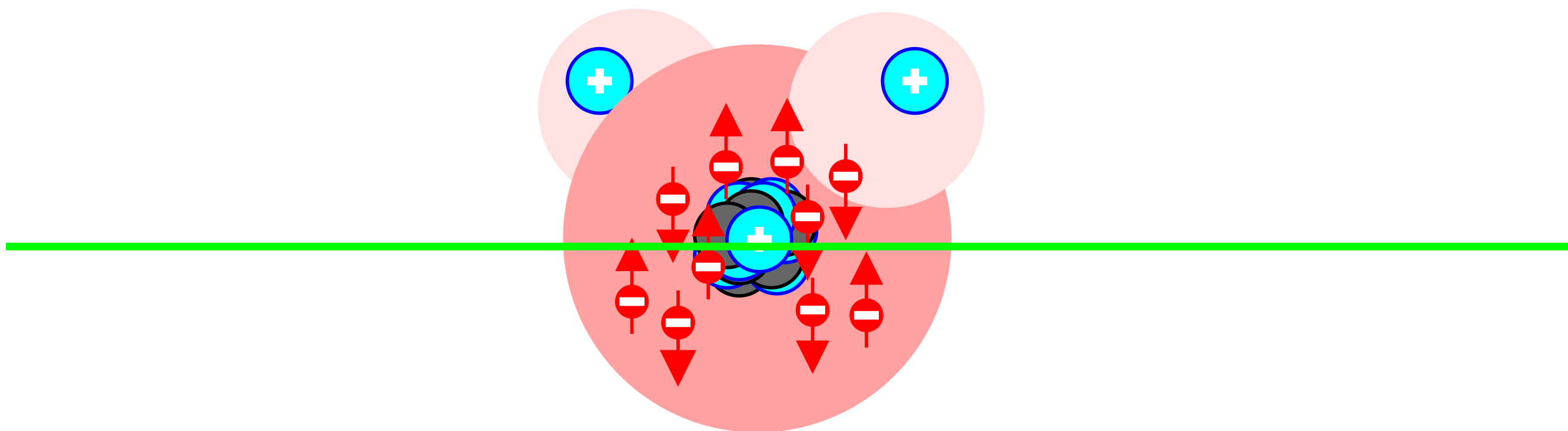
Sharp-edged shadow

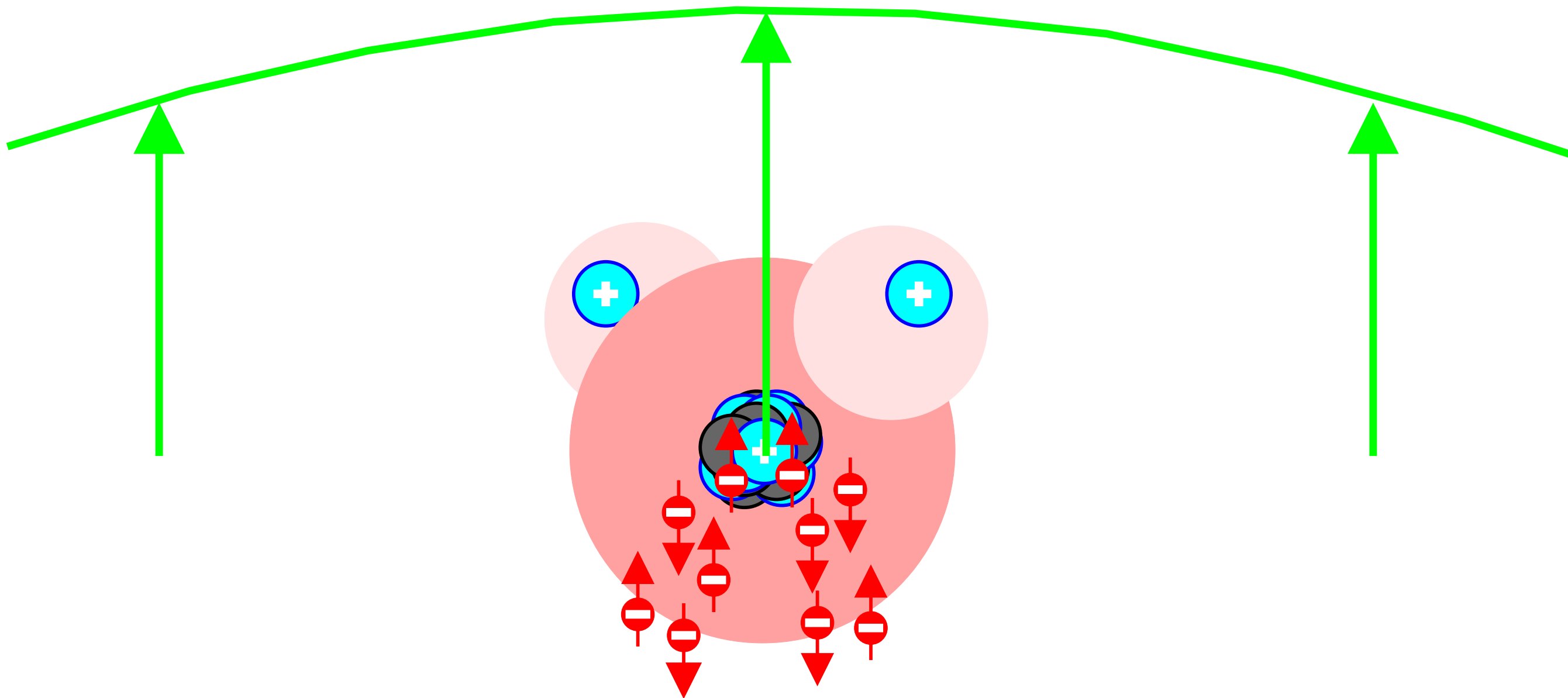


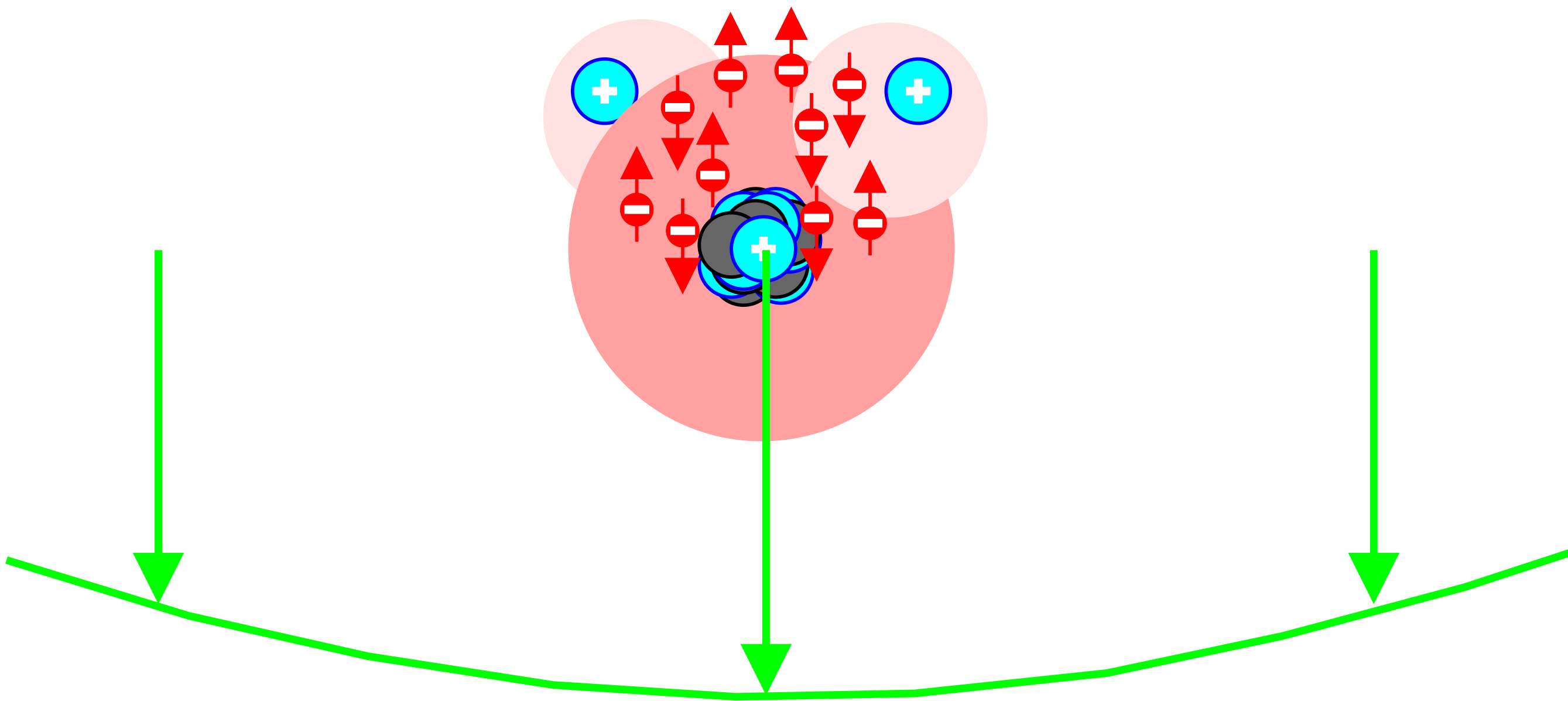
Fuzzy shadow

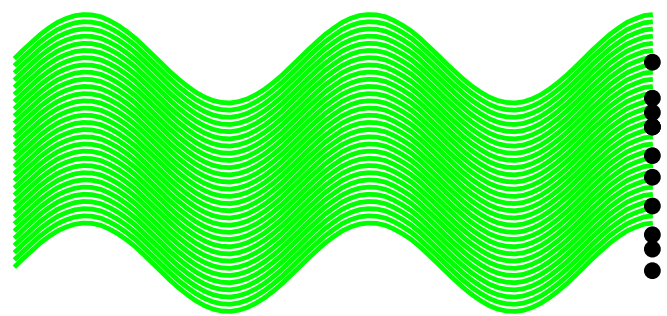


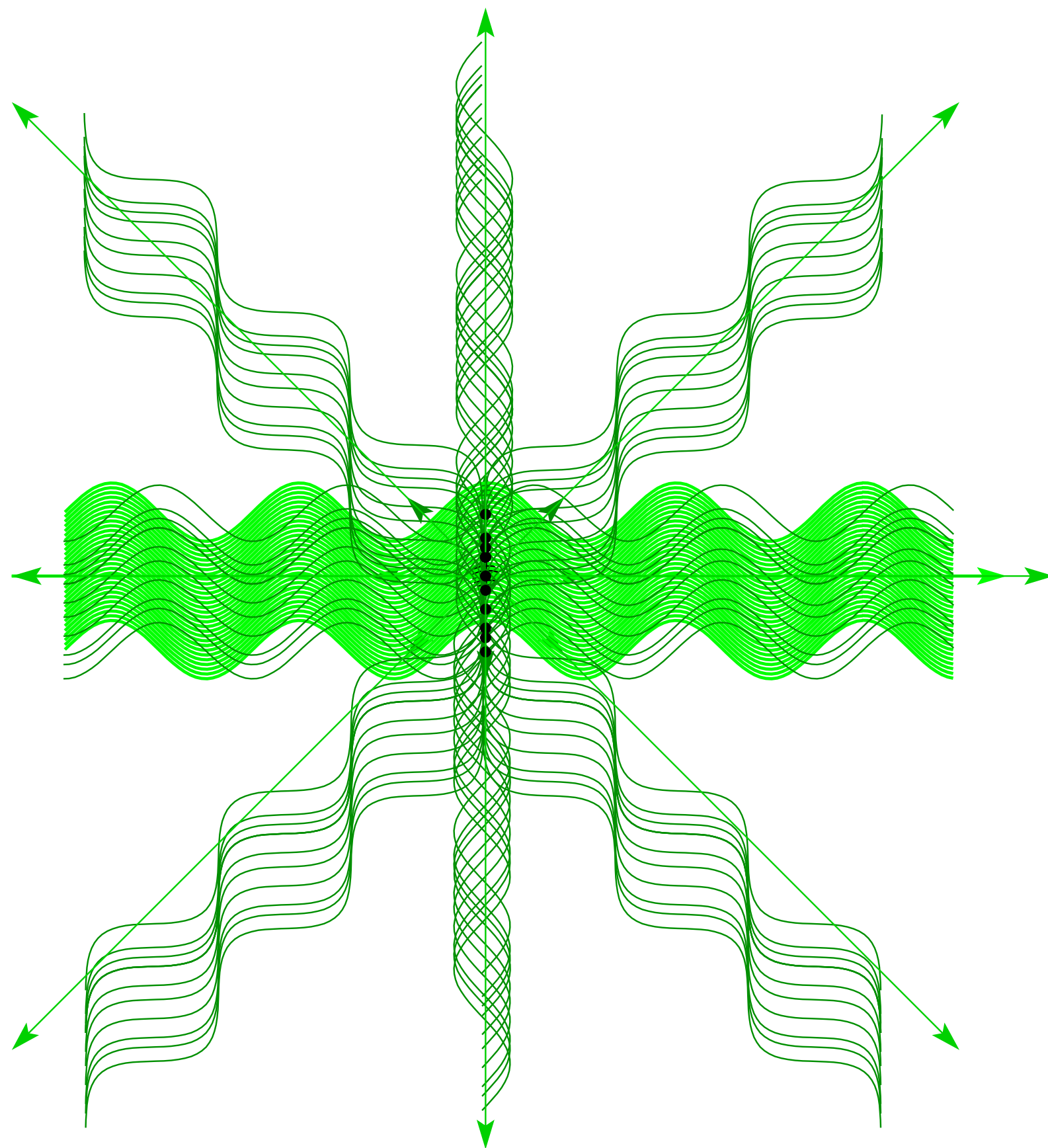


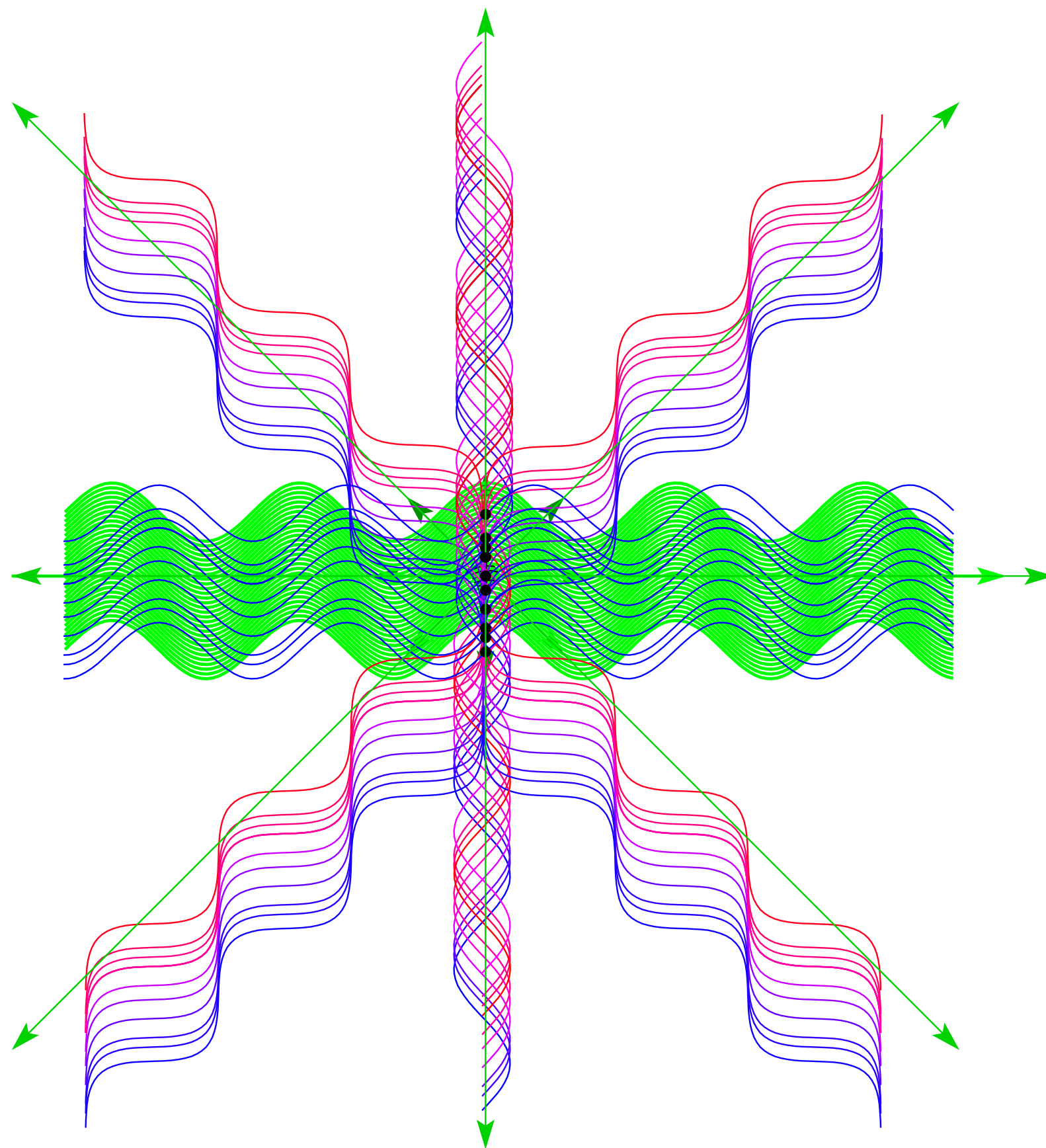


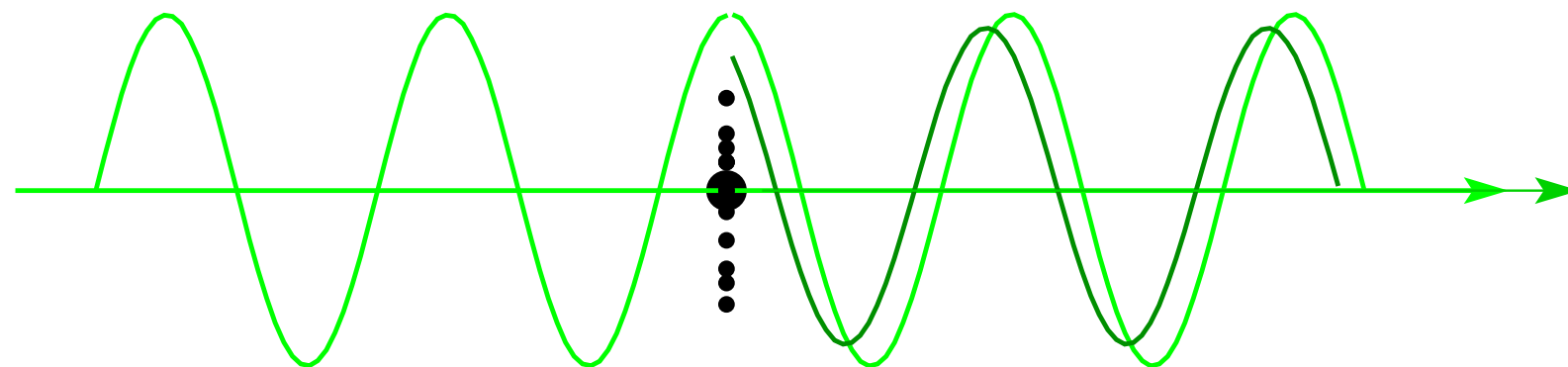


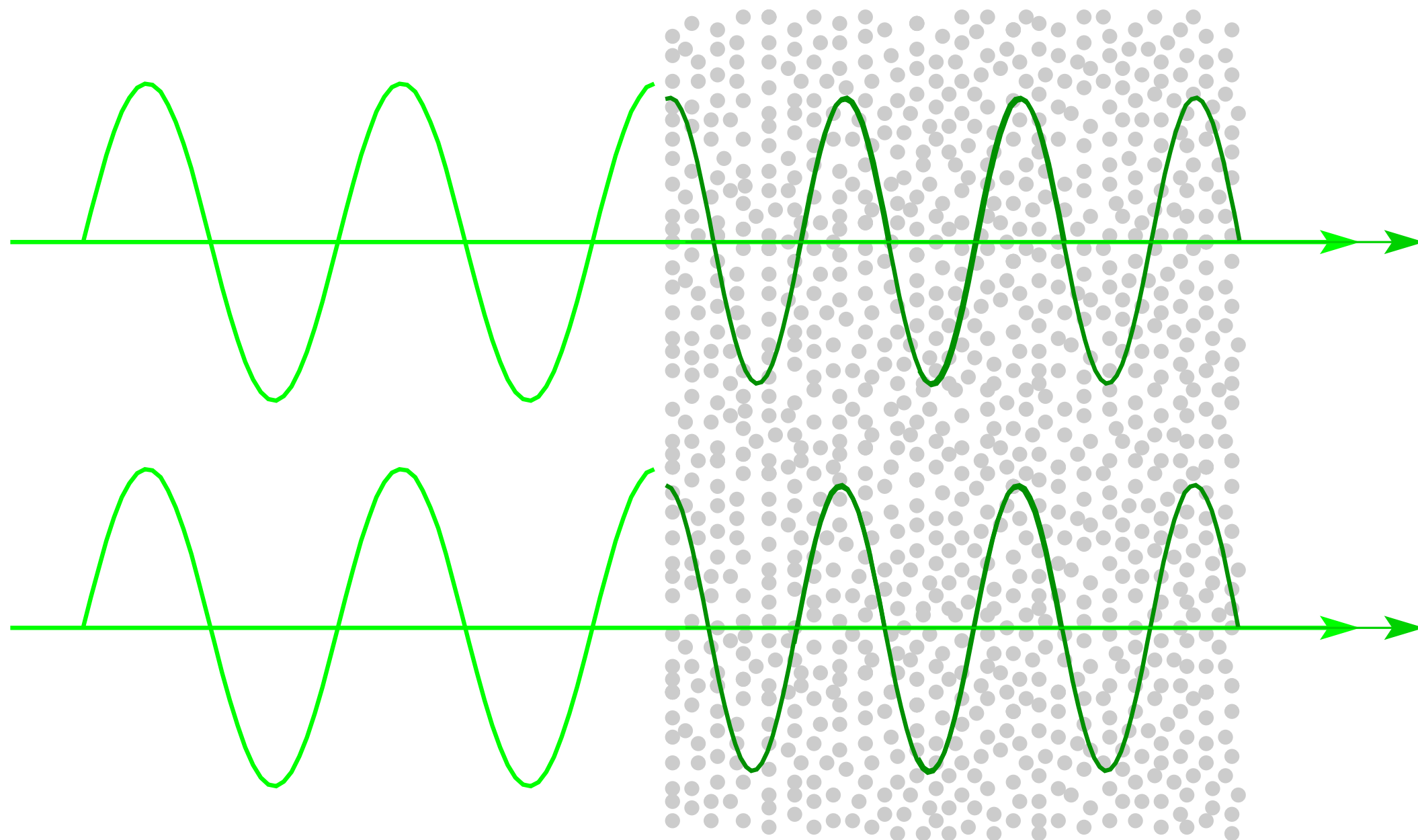


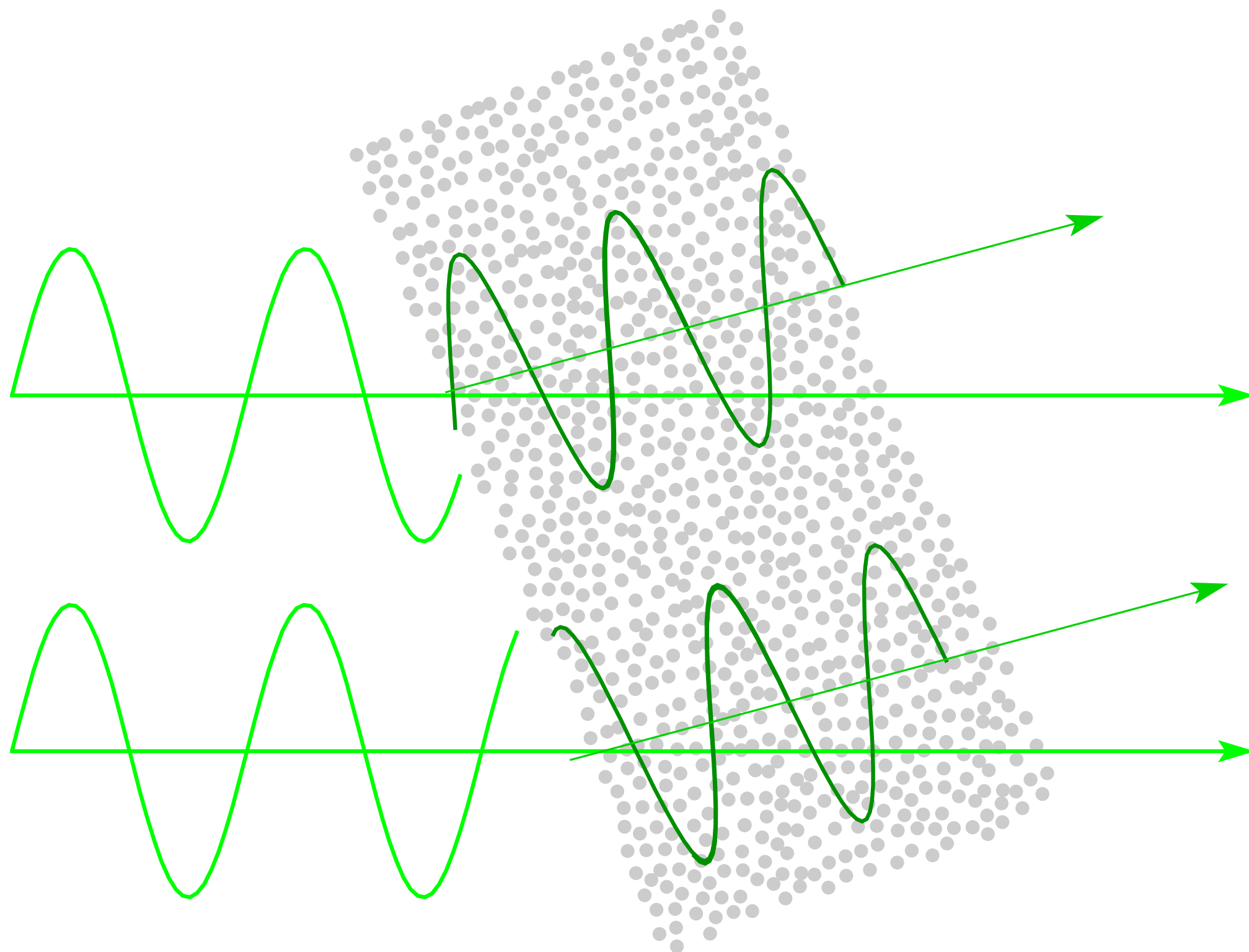


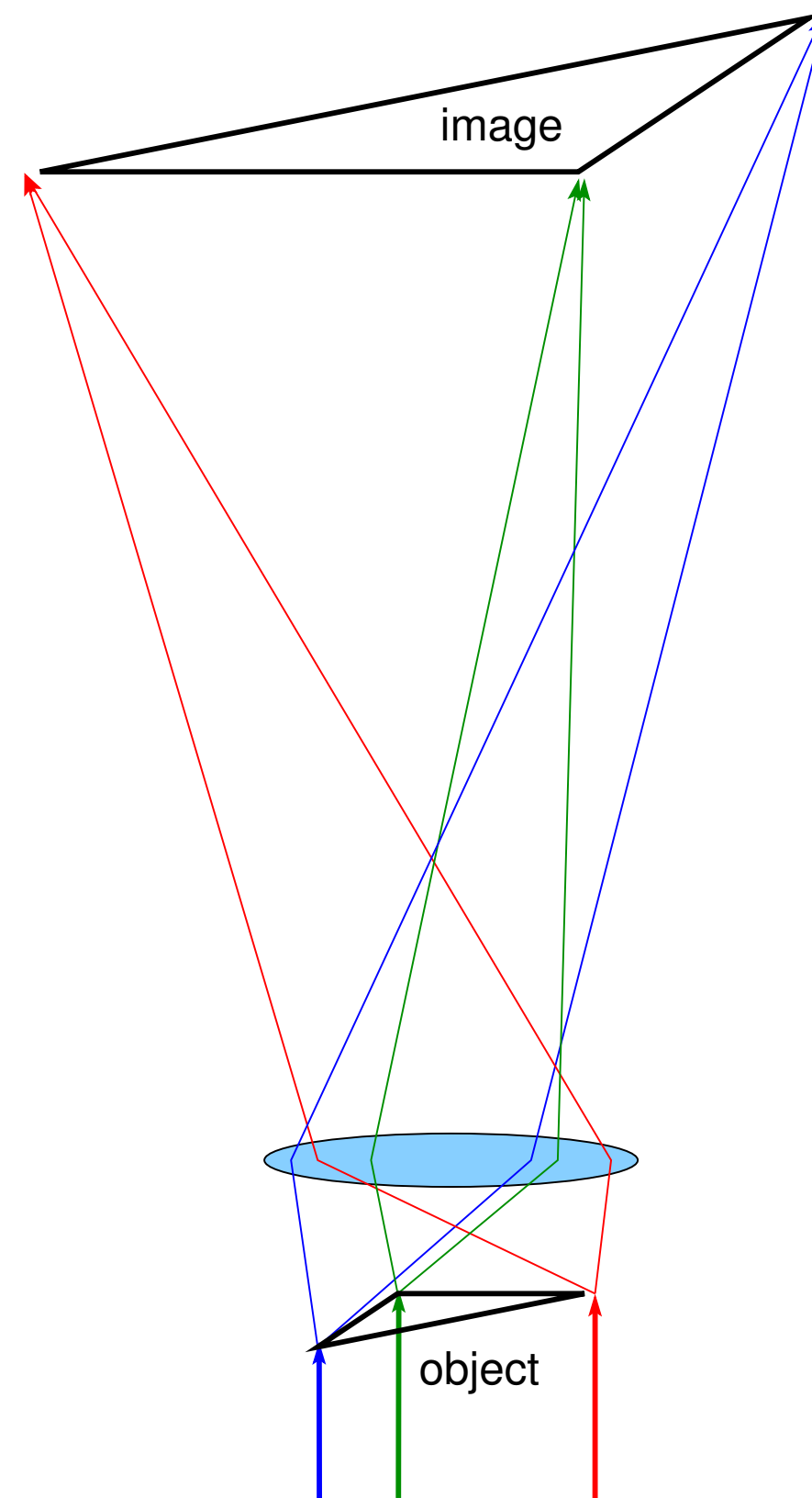
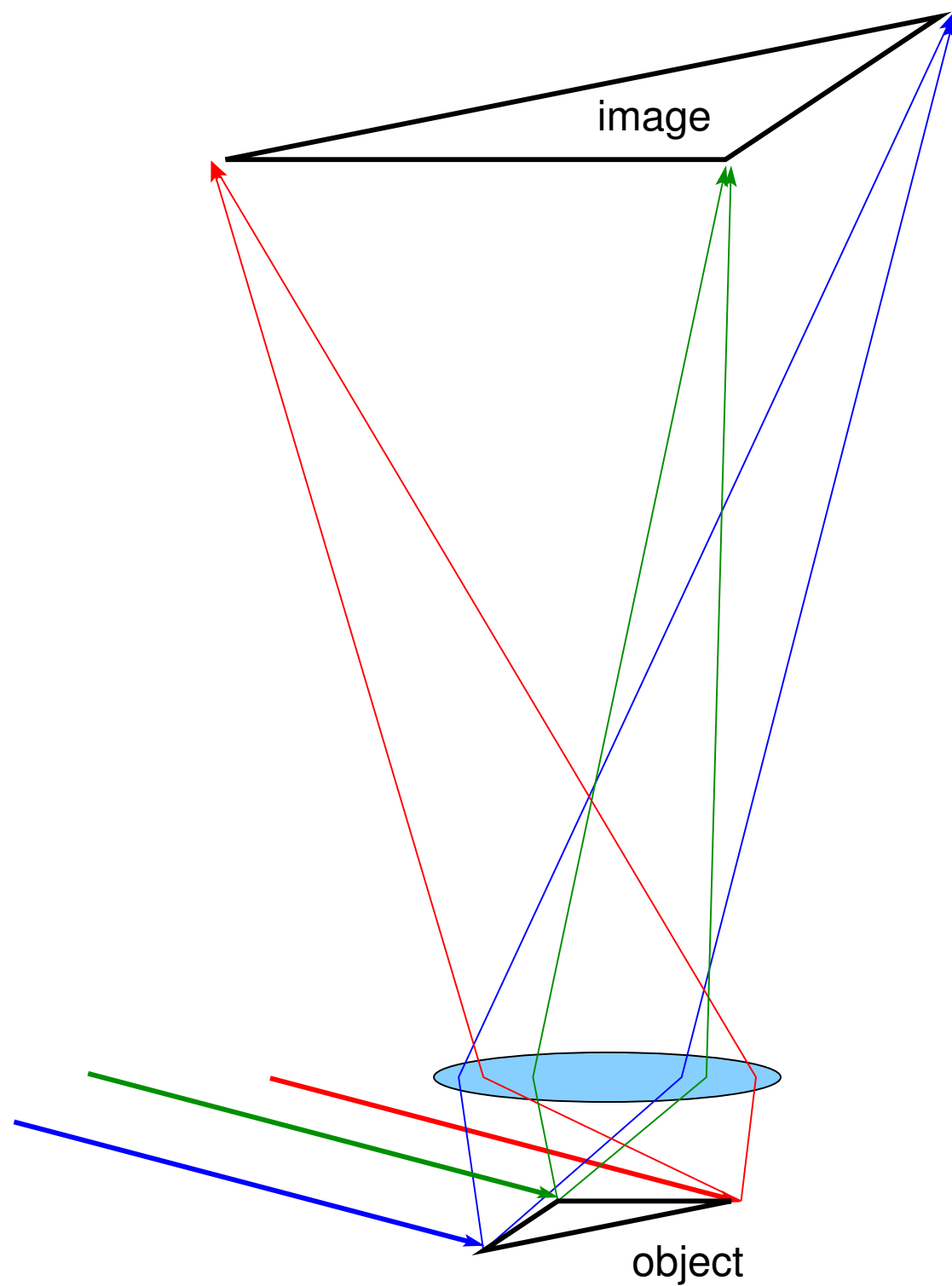


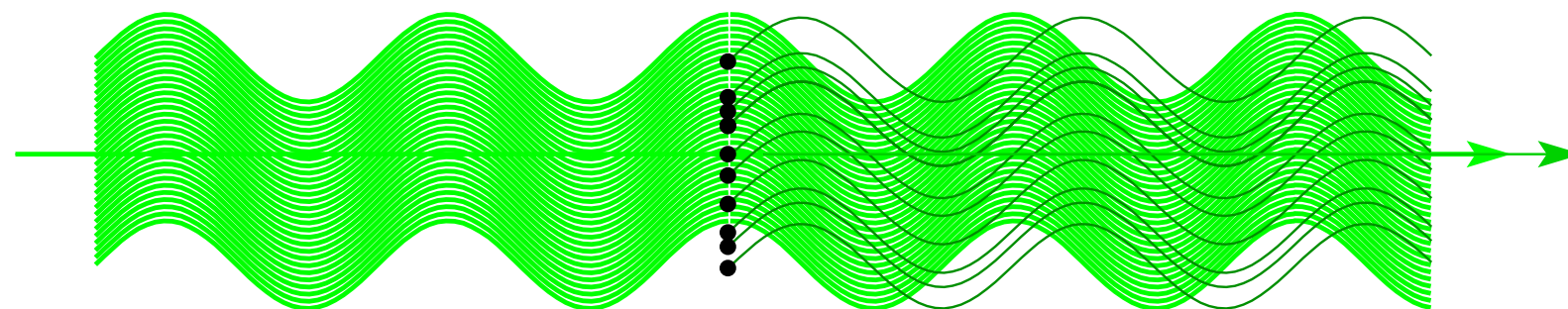


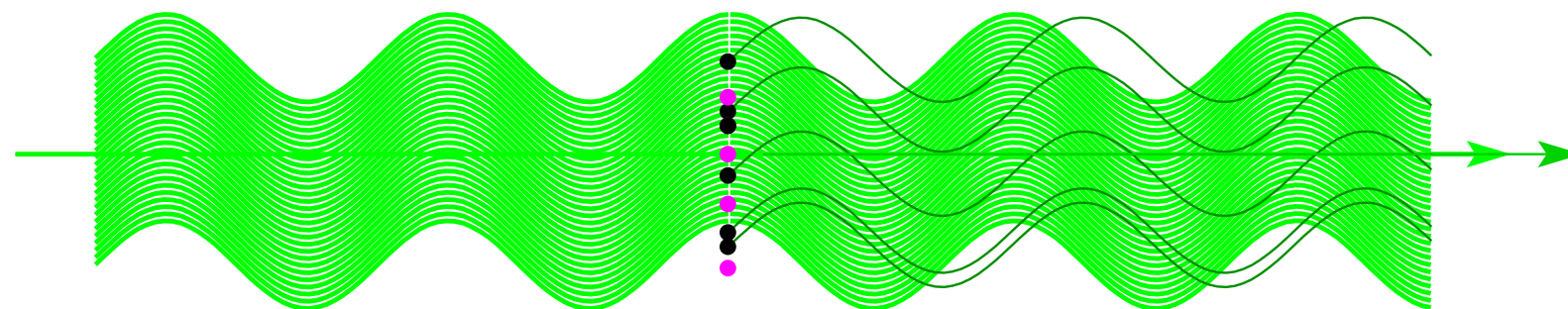


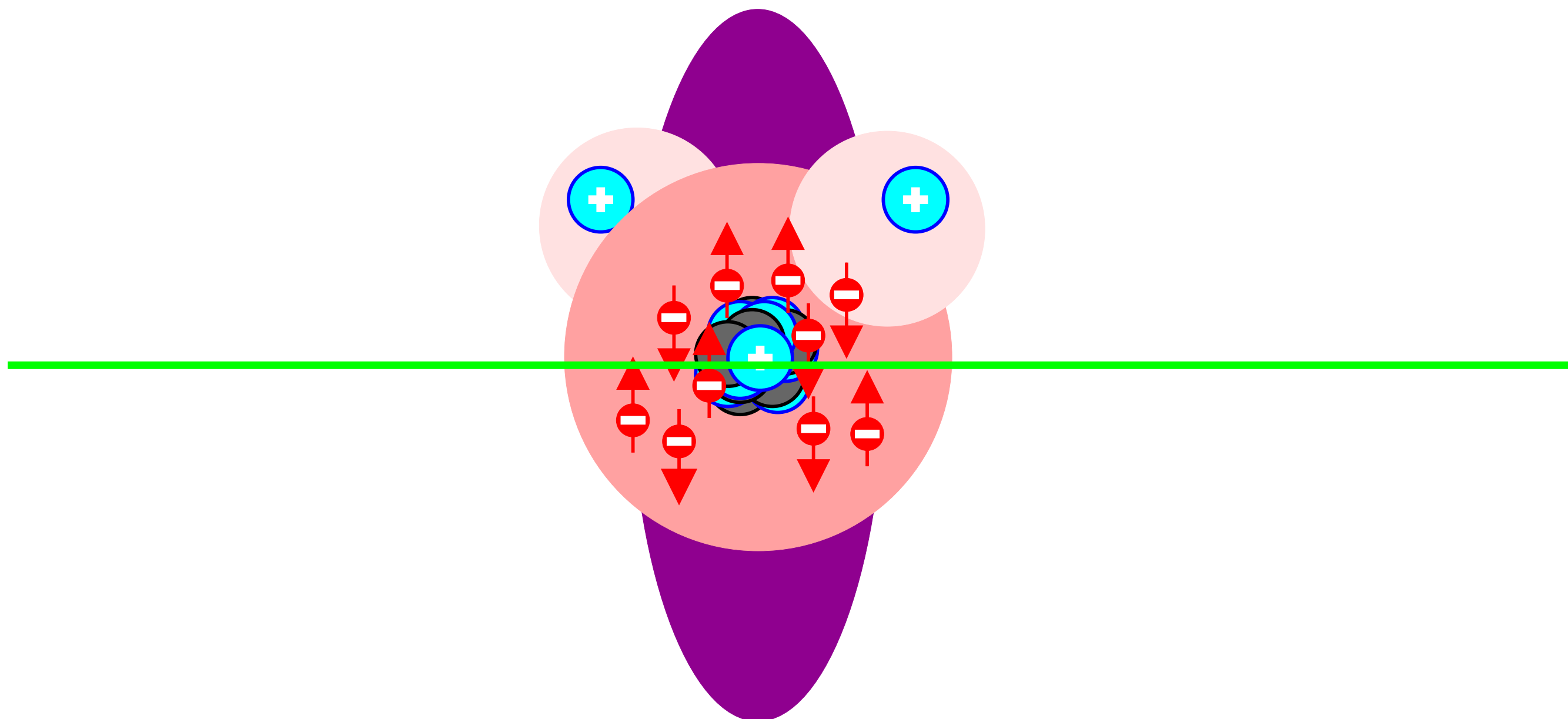


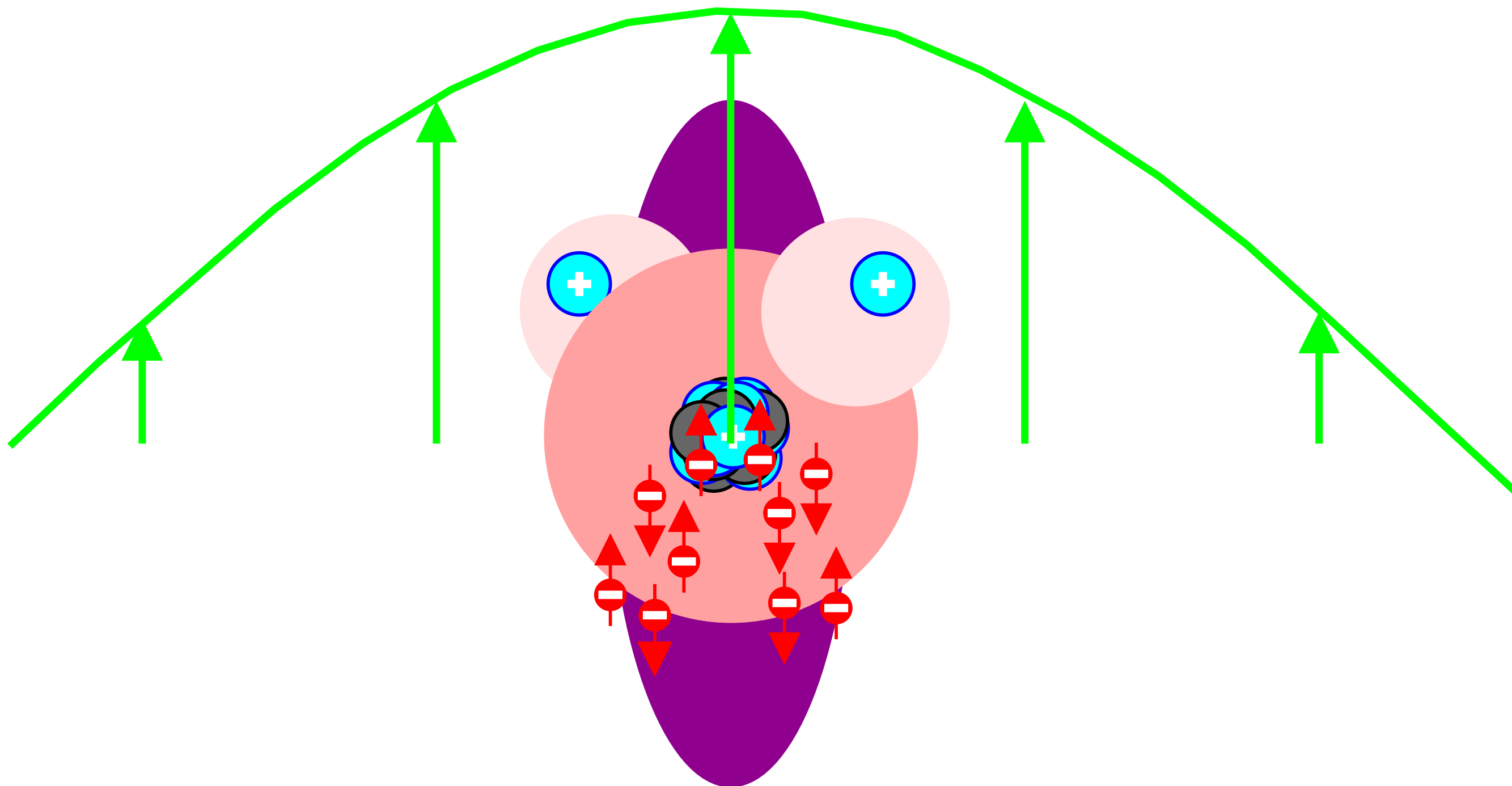


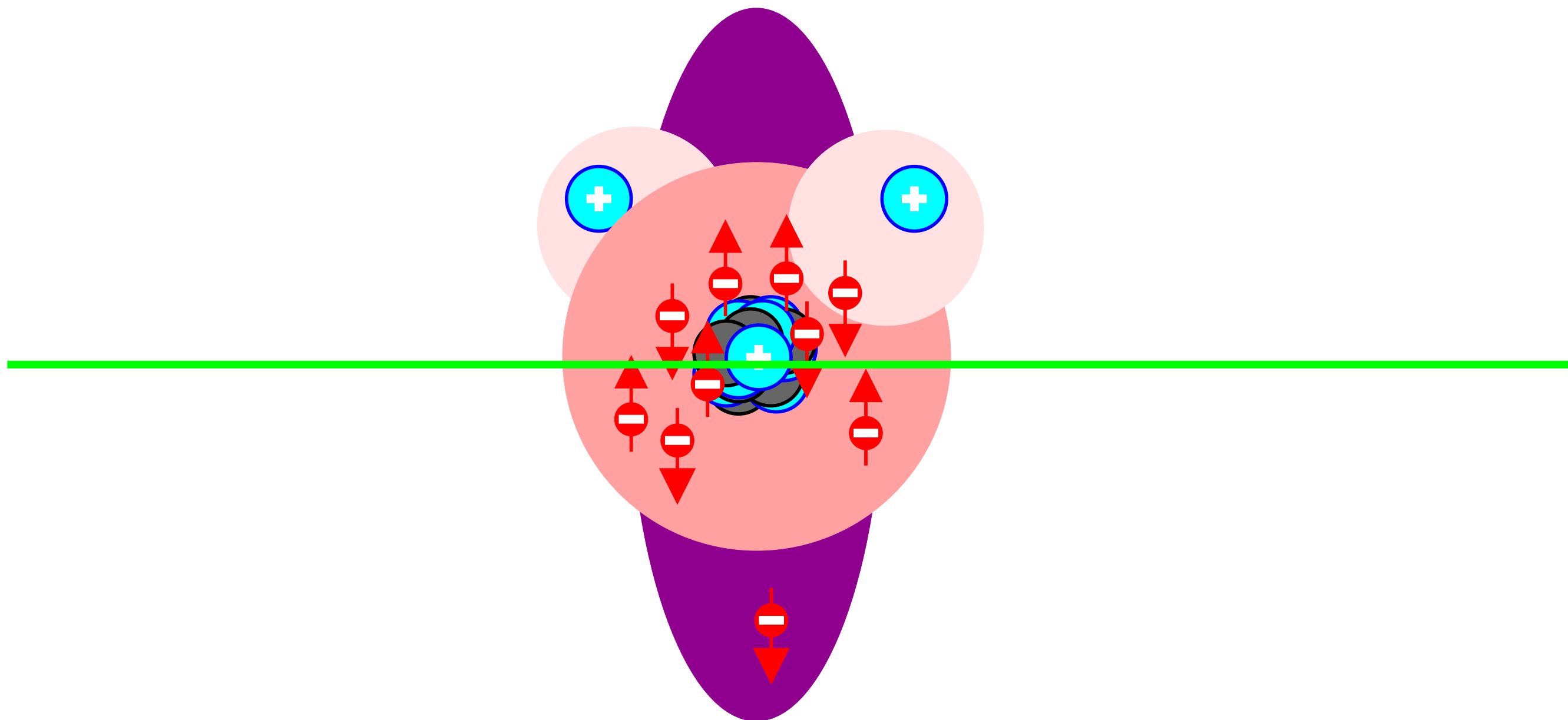


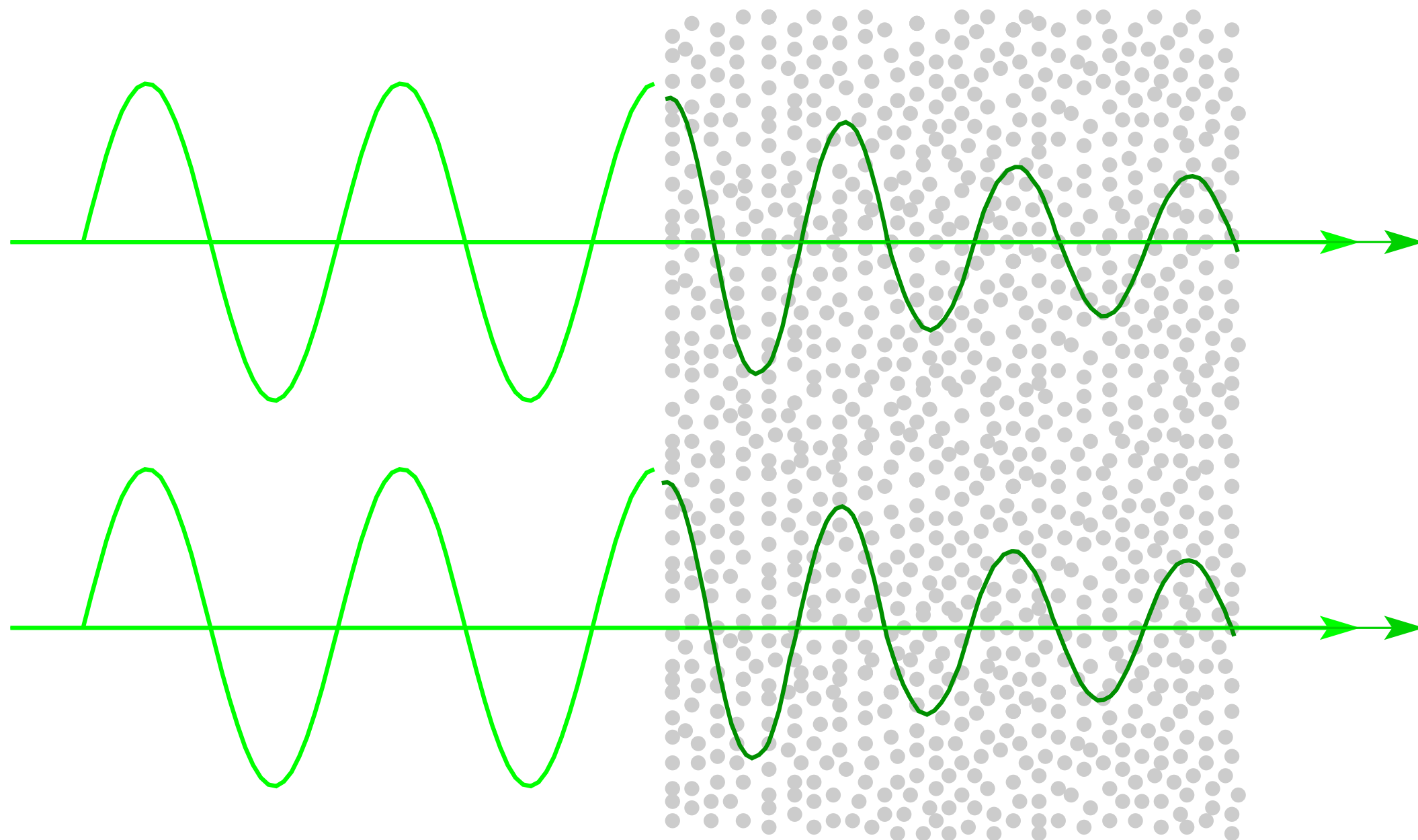


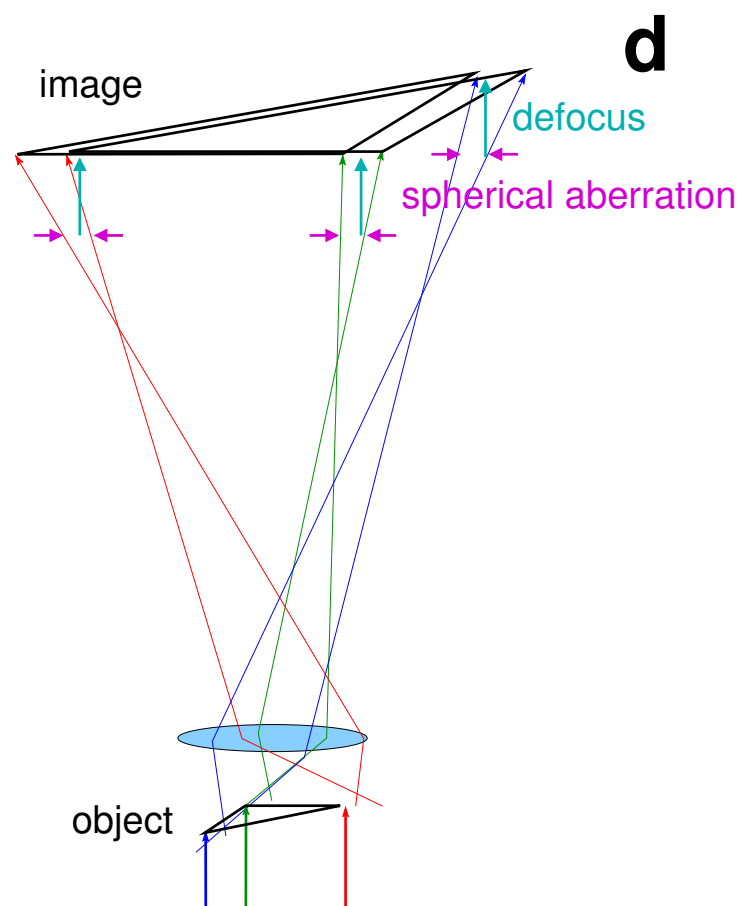
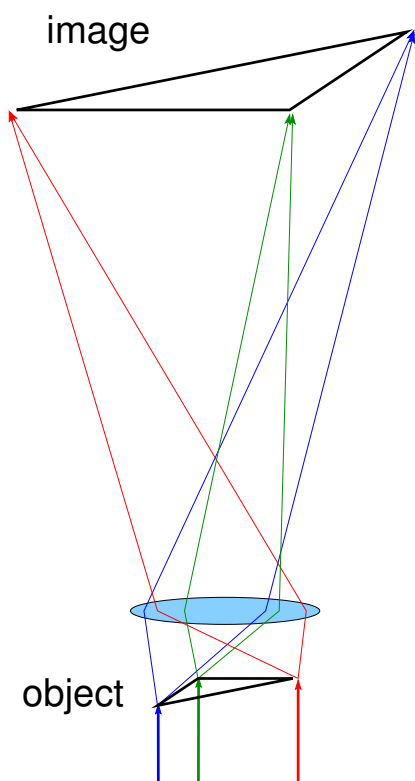
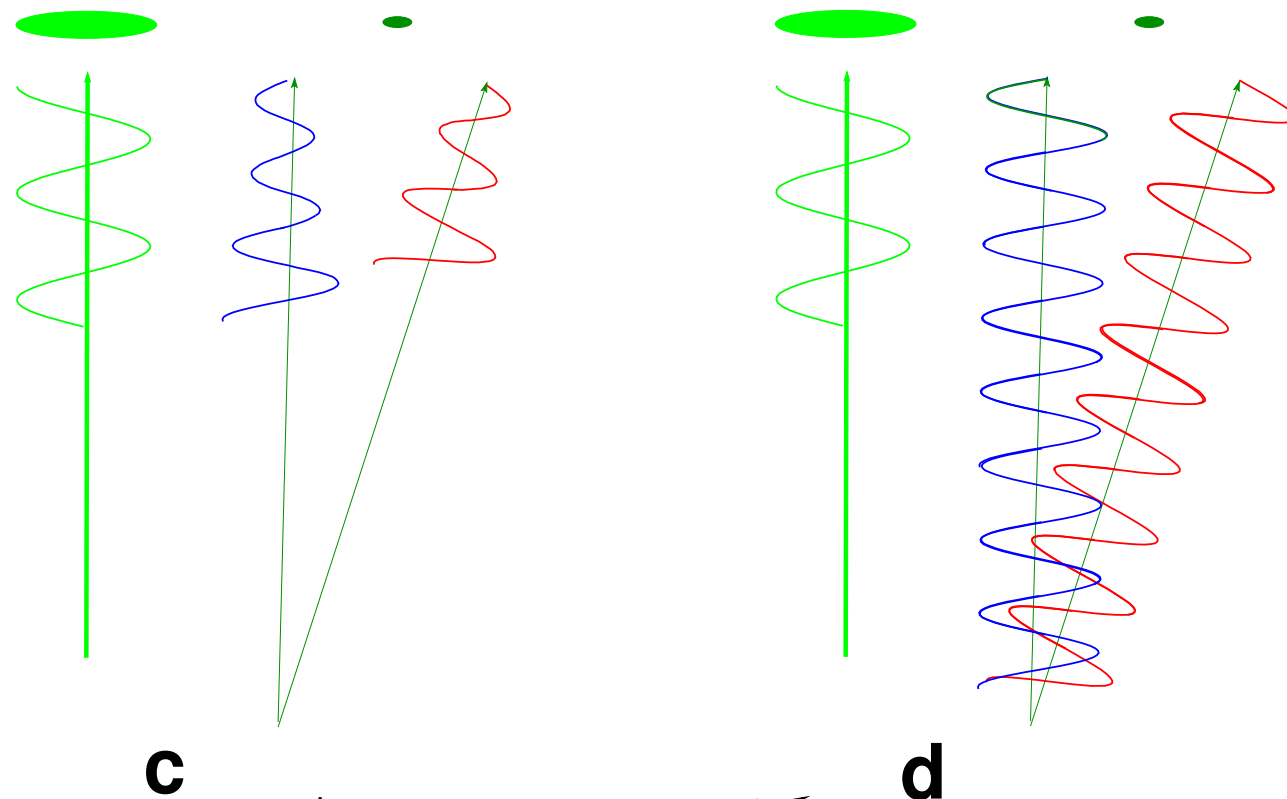
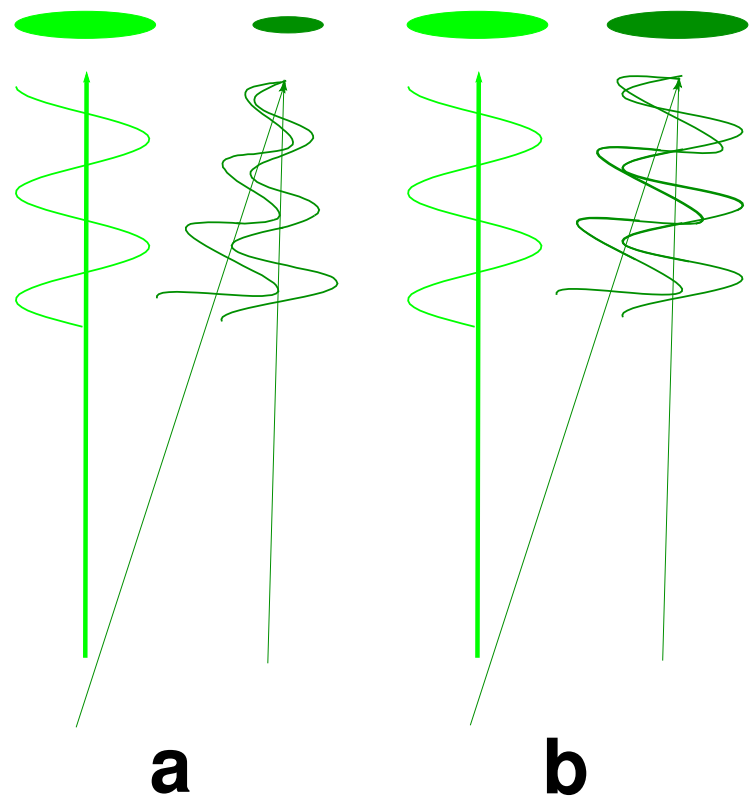


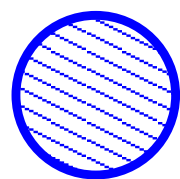




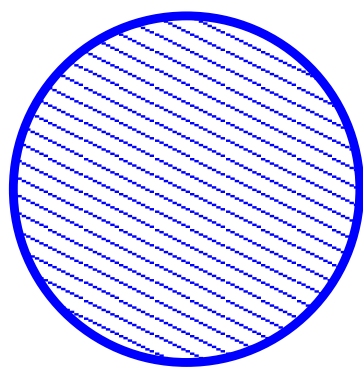




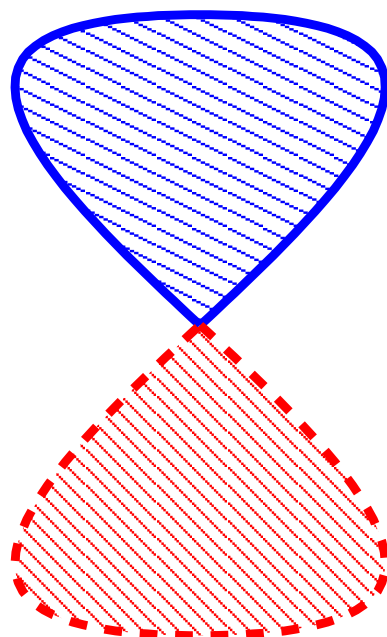




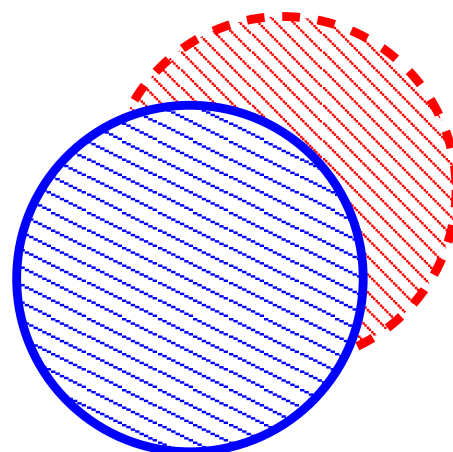
1s



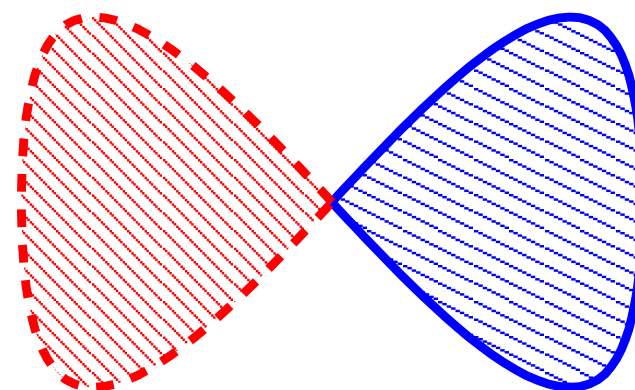
2s



$2p_z$

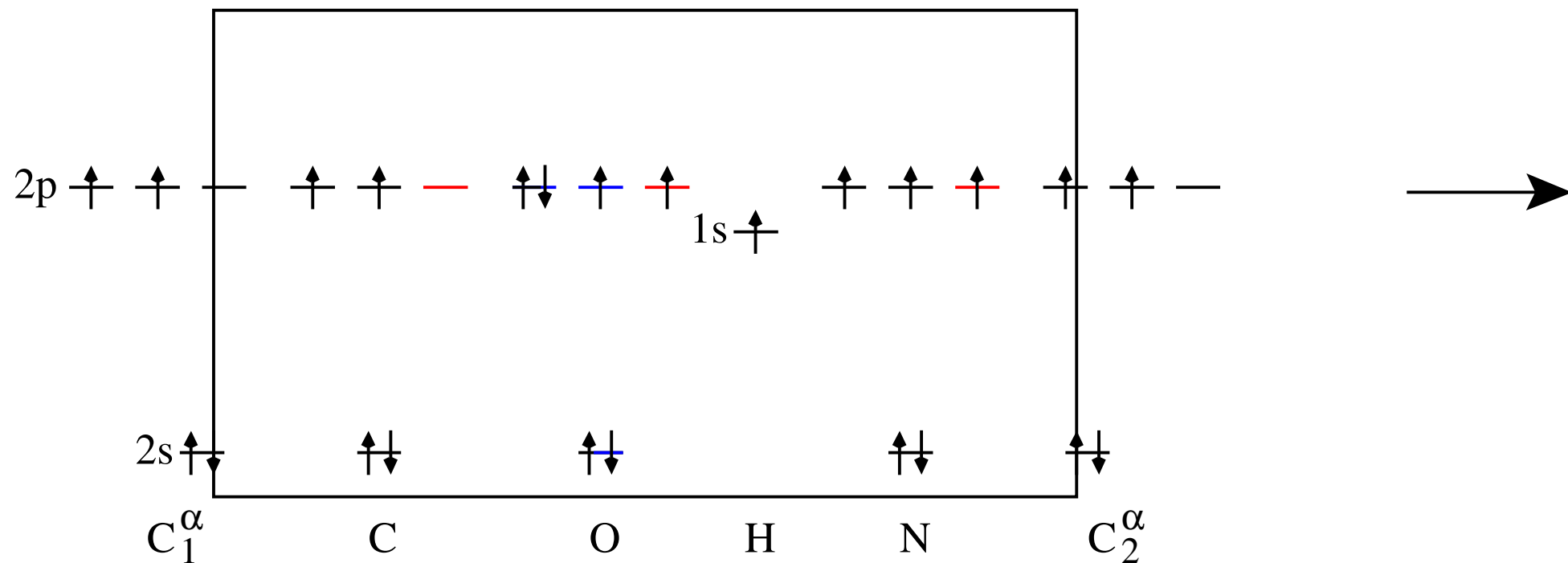


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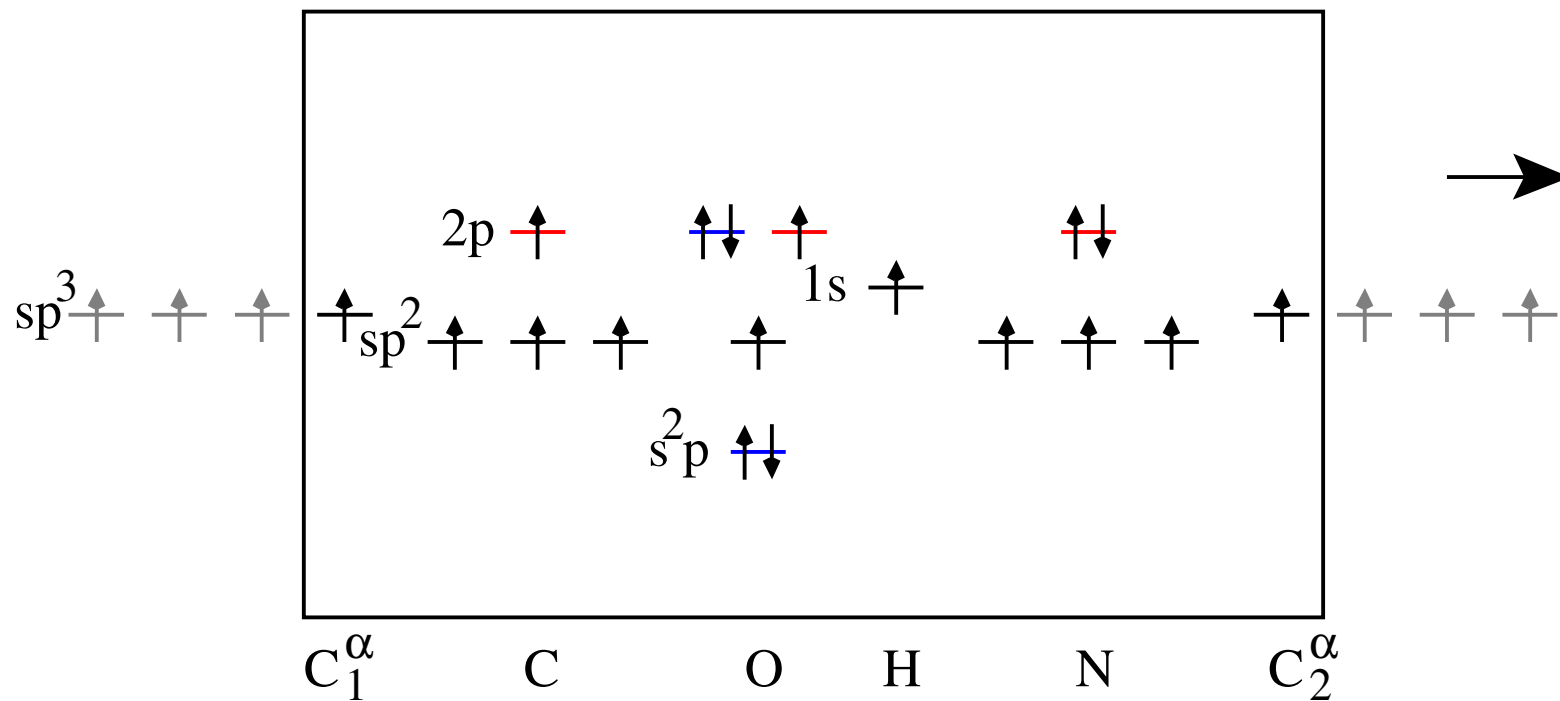


$2p_y$

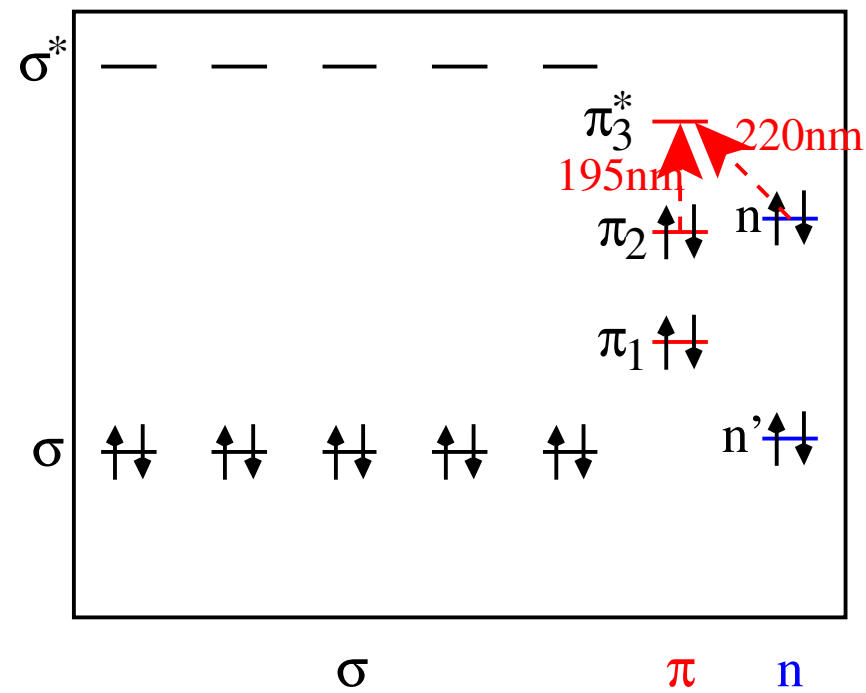
Atomové orbitály

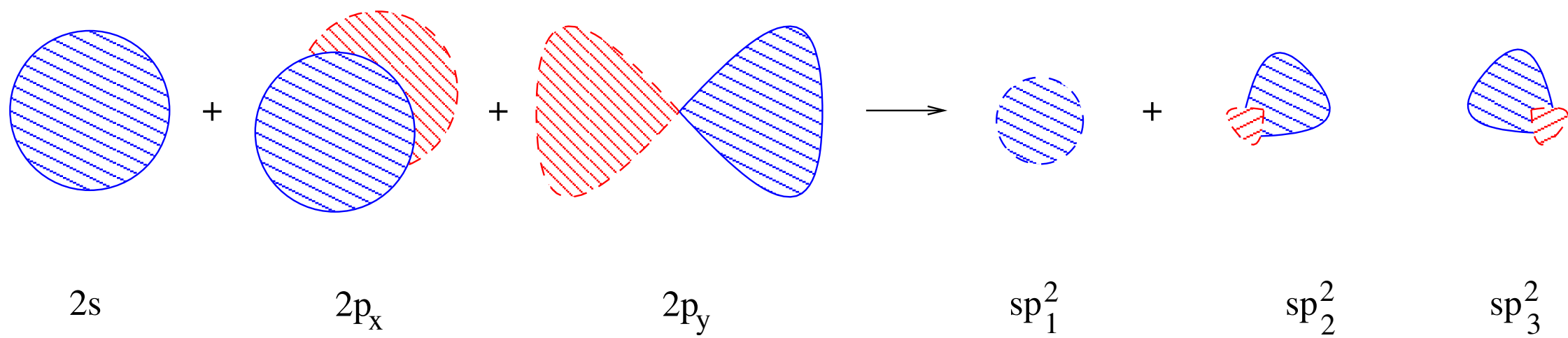
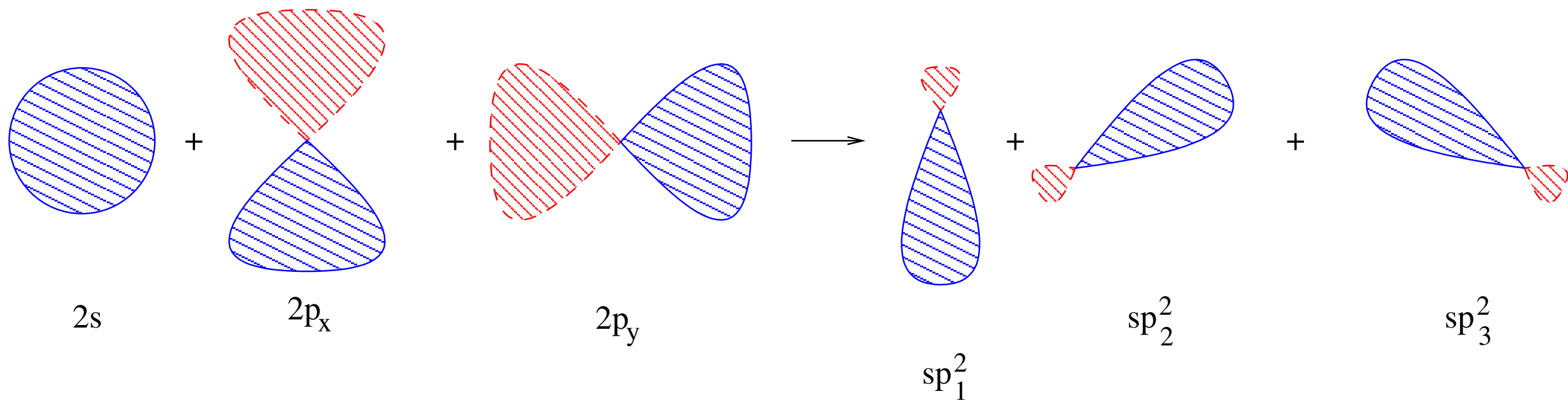


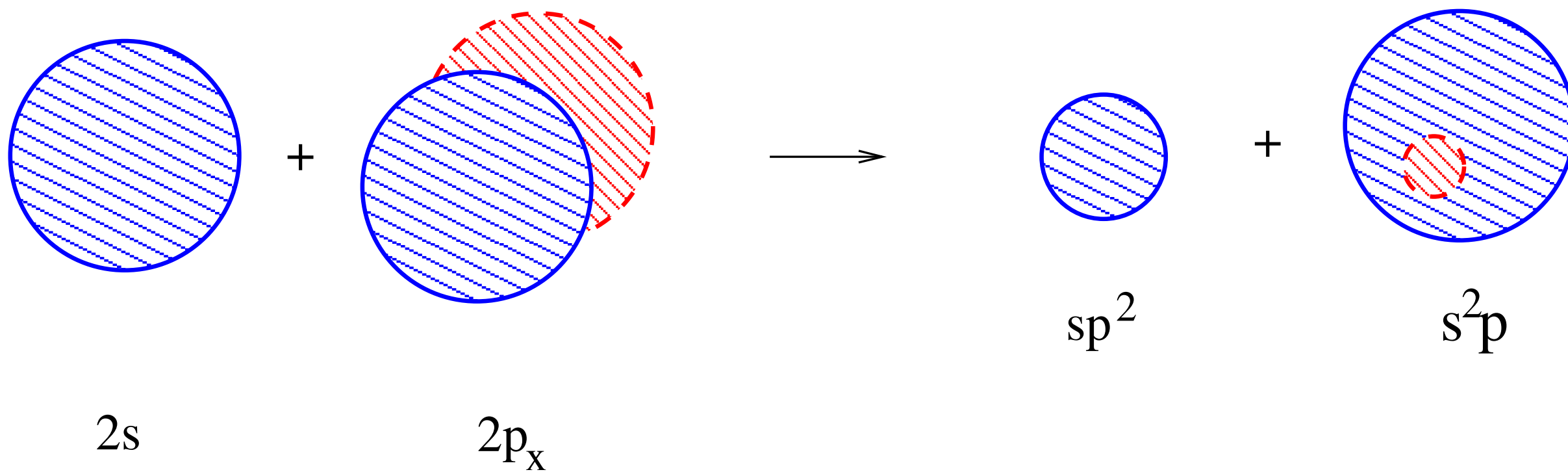
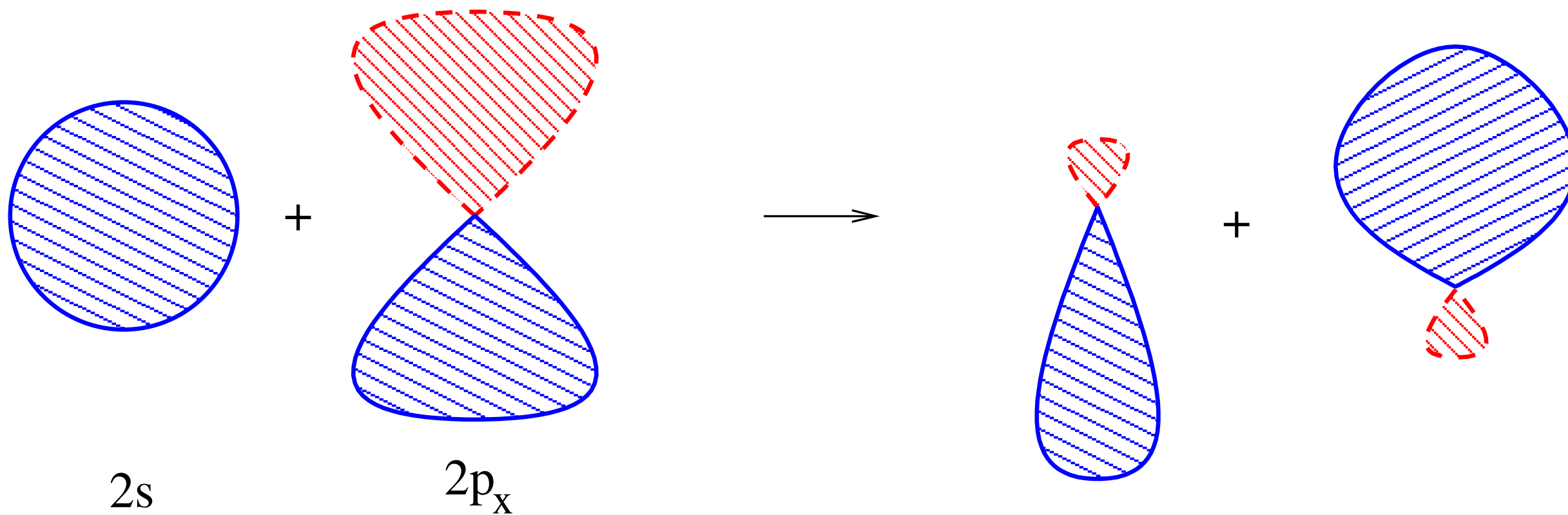
Hybridní orbitály



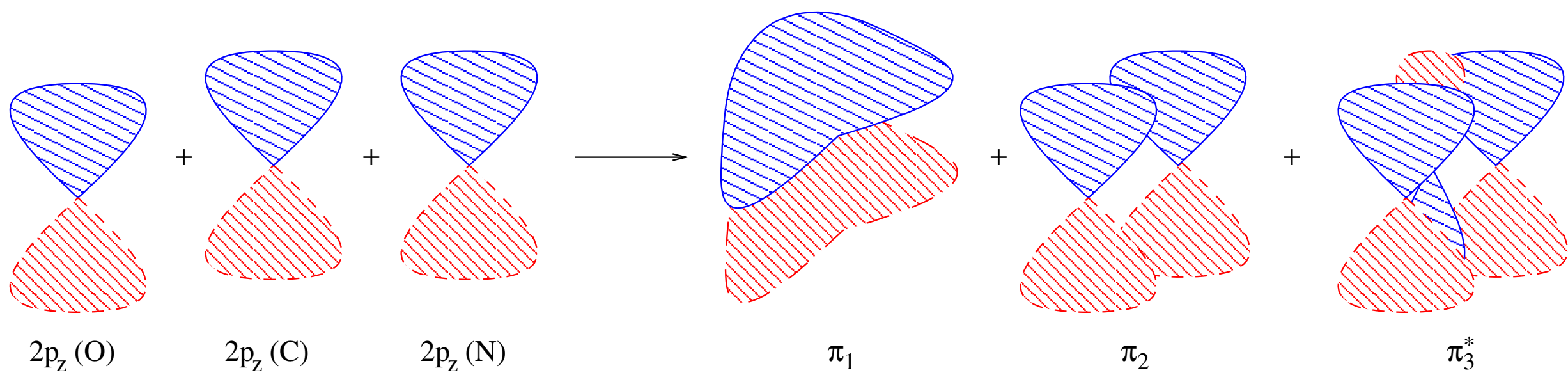
Molekulové orbitály

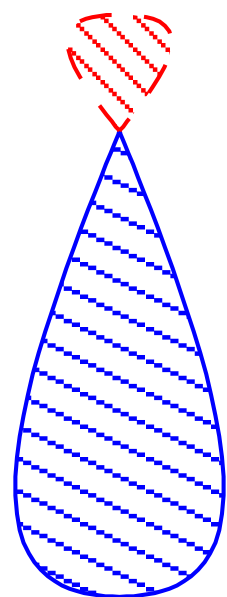




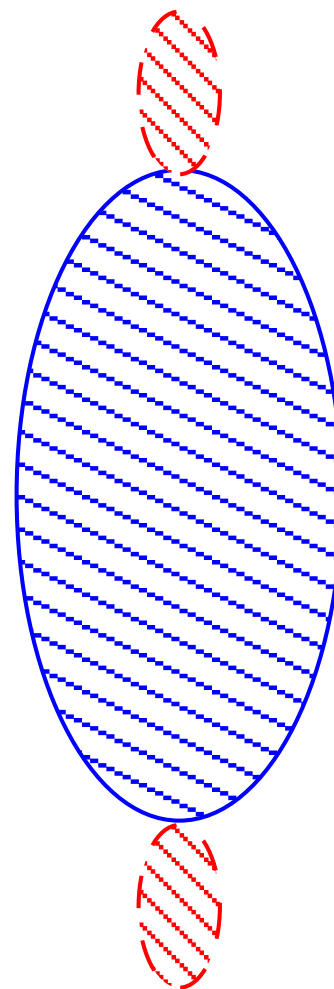
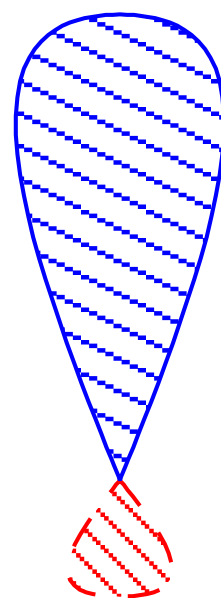


Energy level diagram showing the ground state σ and excited state σ^* . The ground state σ contains five degenerate levels (up and down arrows). The excited state σ^* contains three levels: π_1 , π_2 , and π_3^* . The π state also contains two levels: n and n' . Transitions are indicated by red arrows: a solid arrow from π_2 to π_3^* labeled 195nm, and a dashed arrow from n to π_3^* labeled 220nm.





+



+



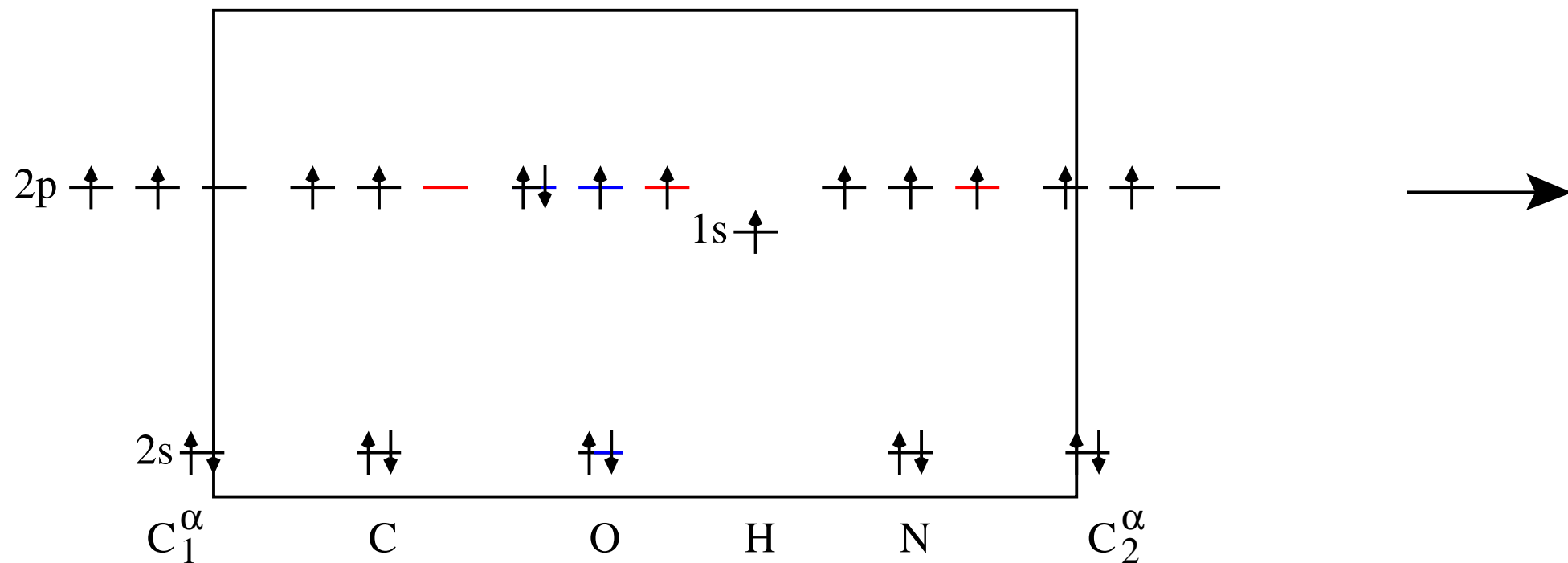
$sp^2(X)$

$sp^2(Y)$

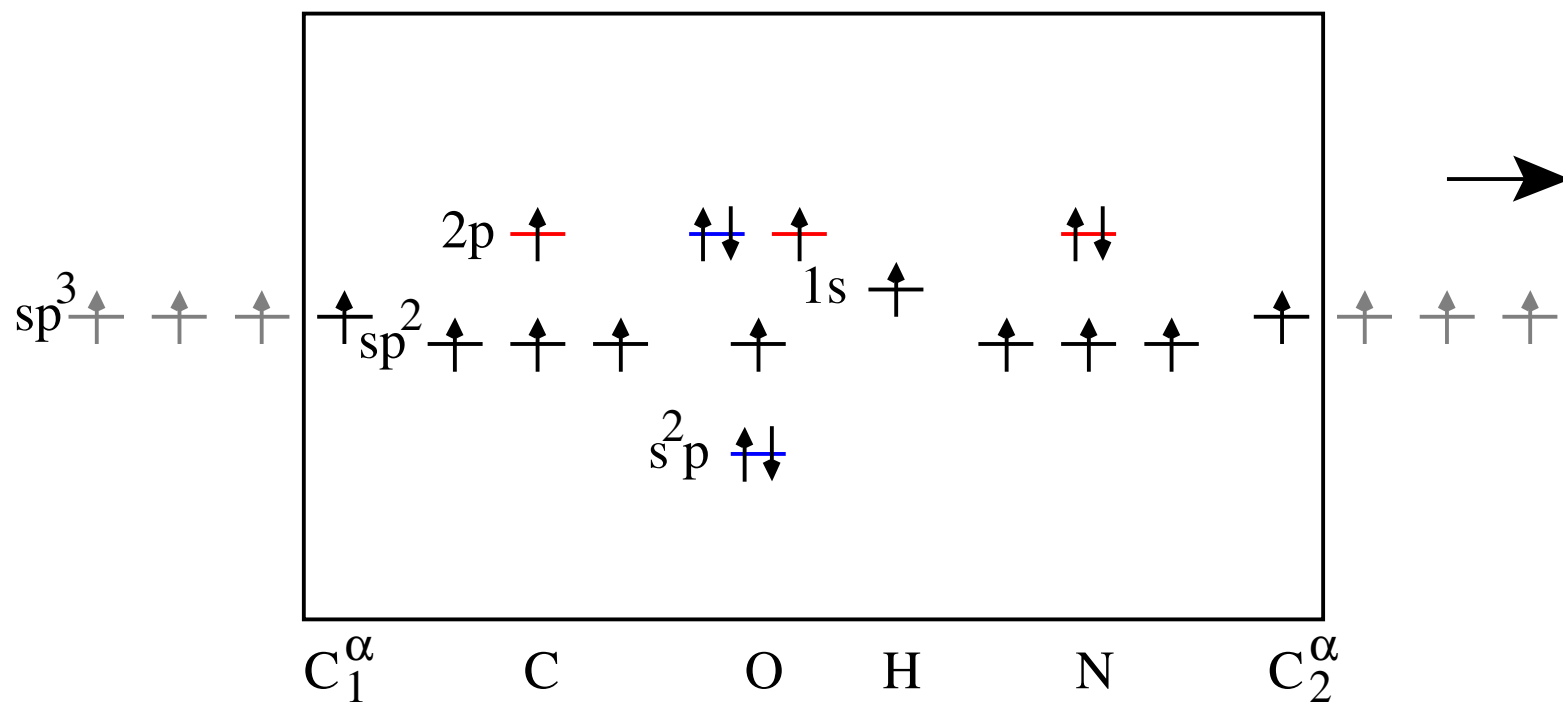
$\sigma(XY)$

$\sigma^*(XY)$

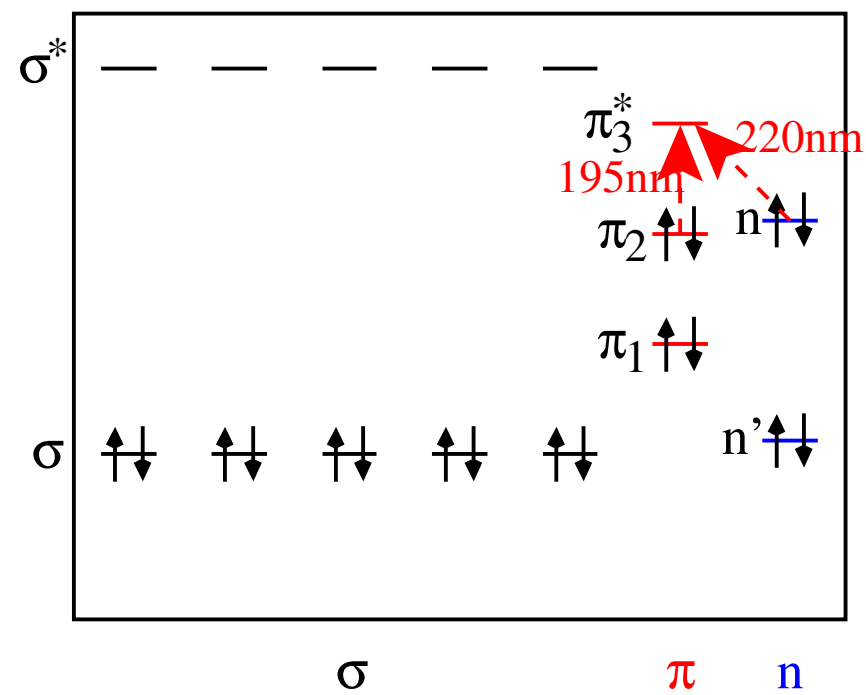
Atomové orbitály

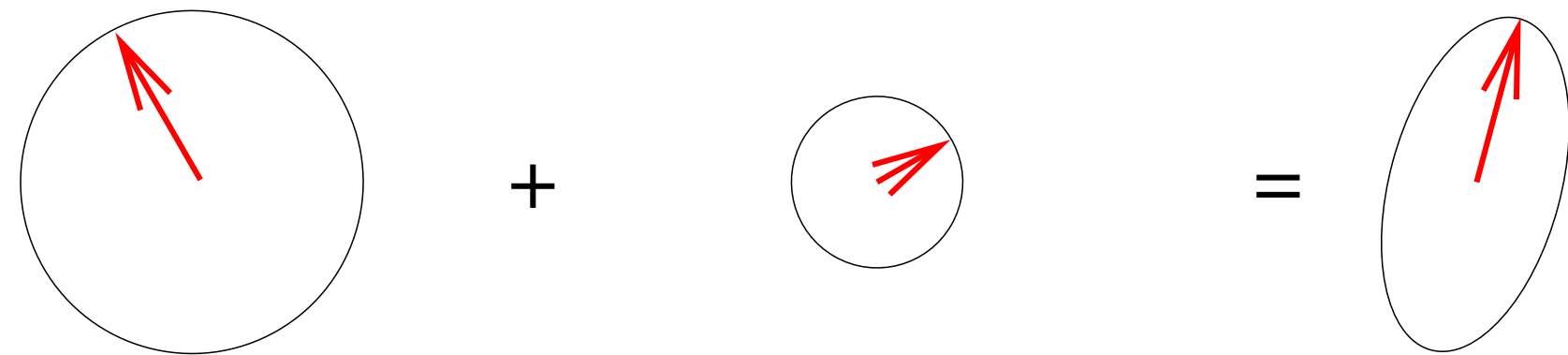
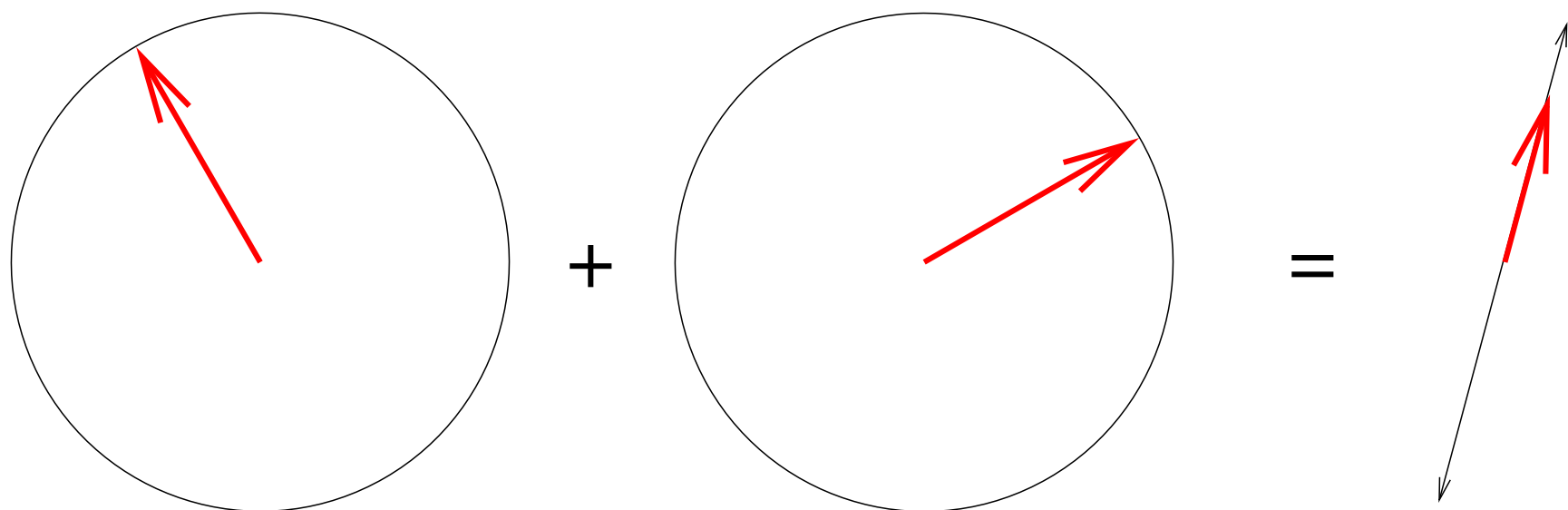
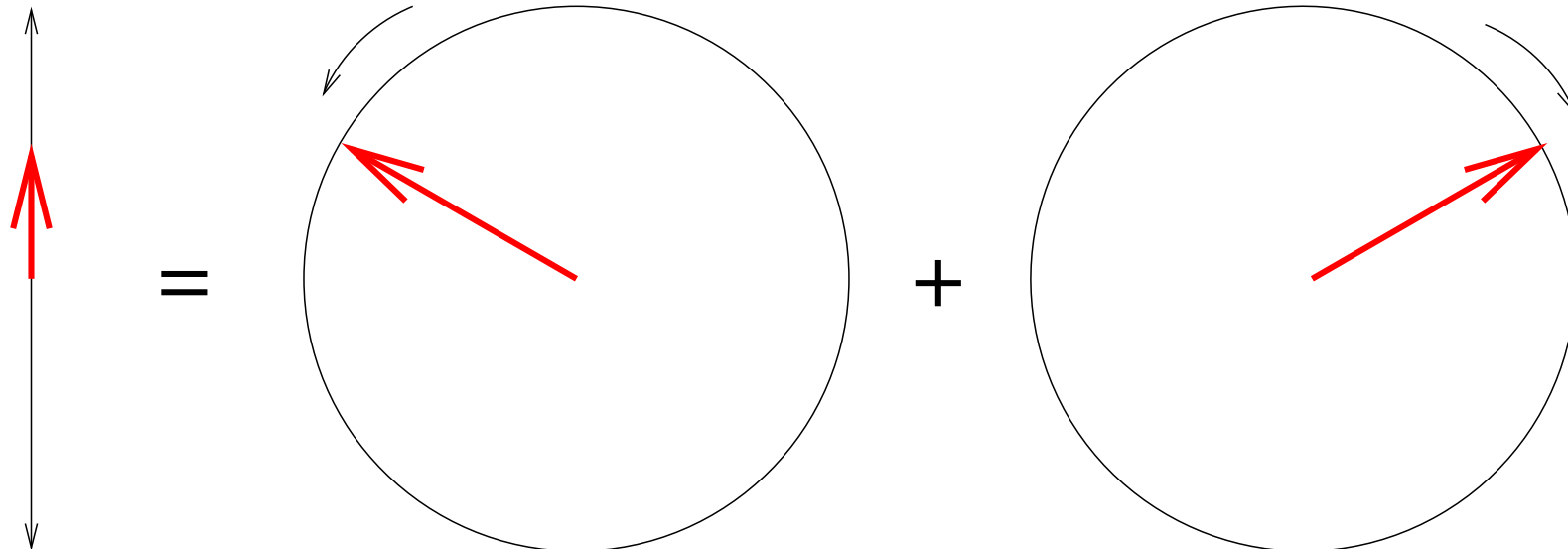


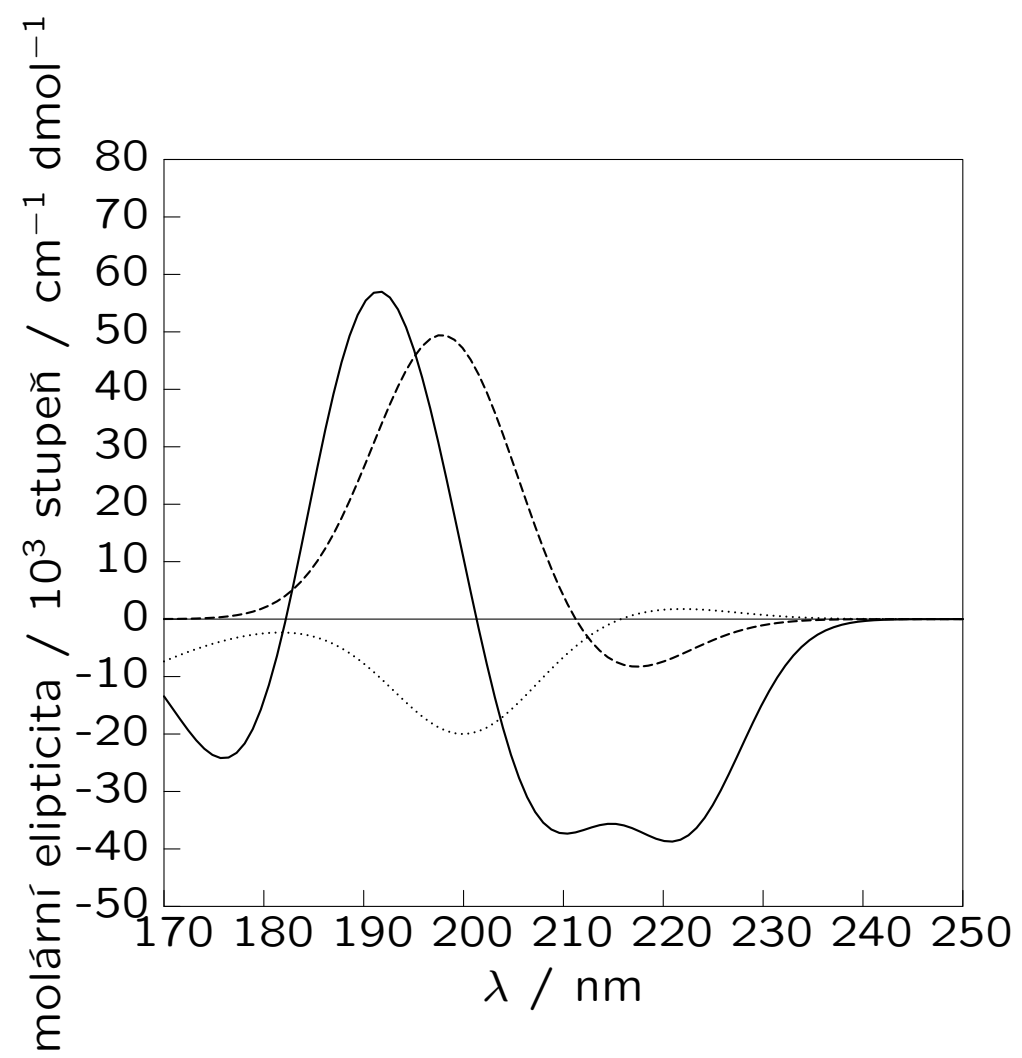
Hybridní orbitály

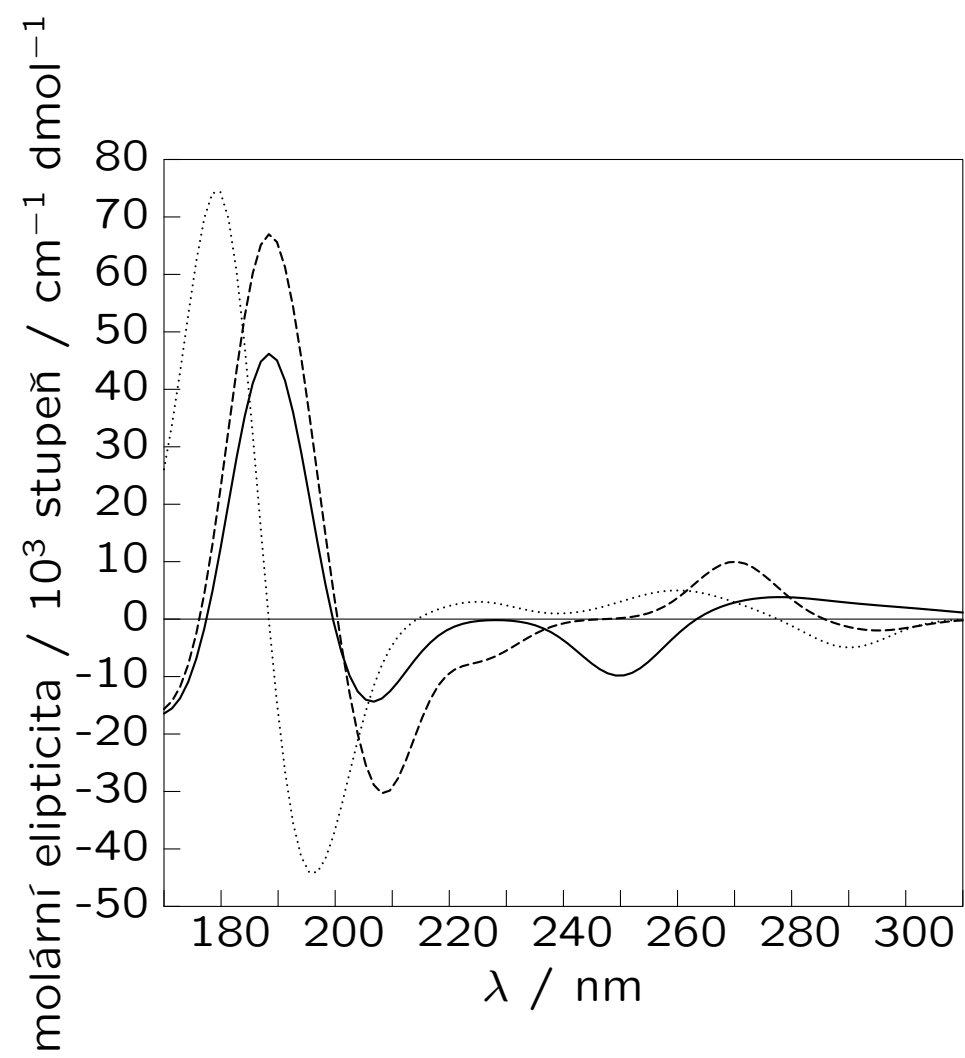


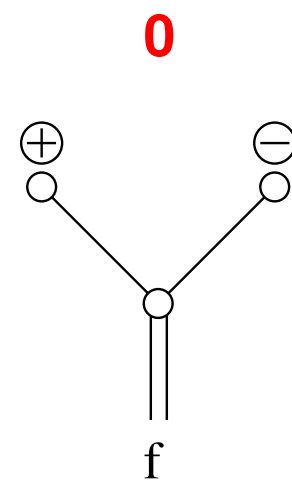
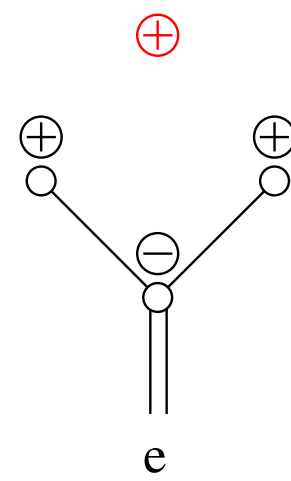
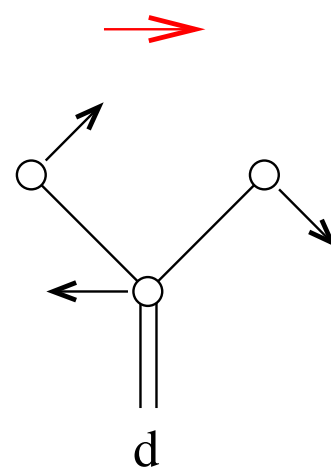
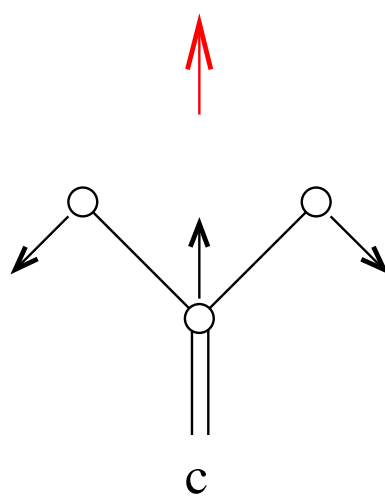
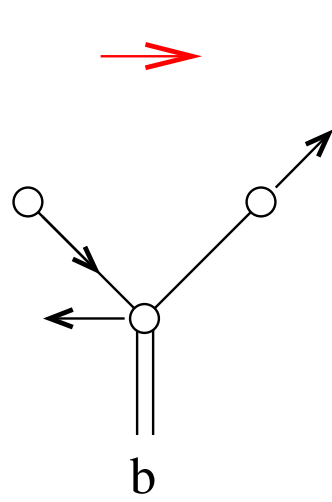
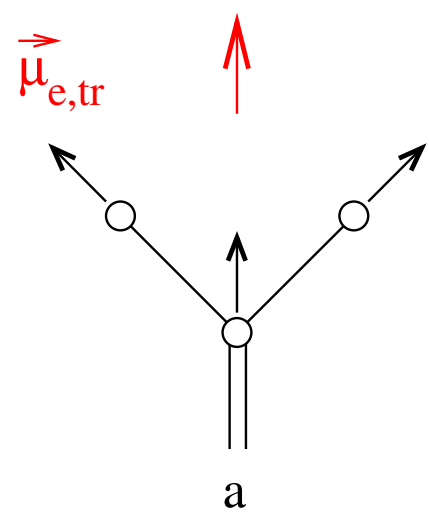
Molekulové orbitály

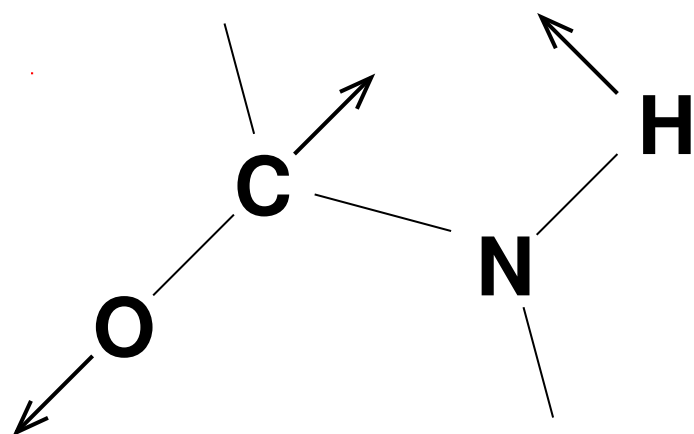




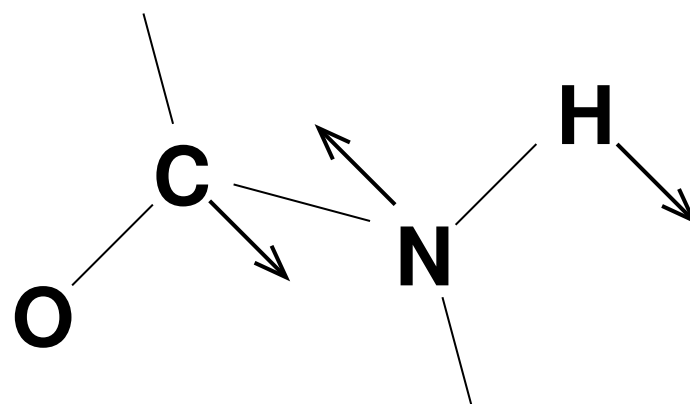




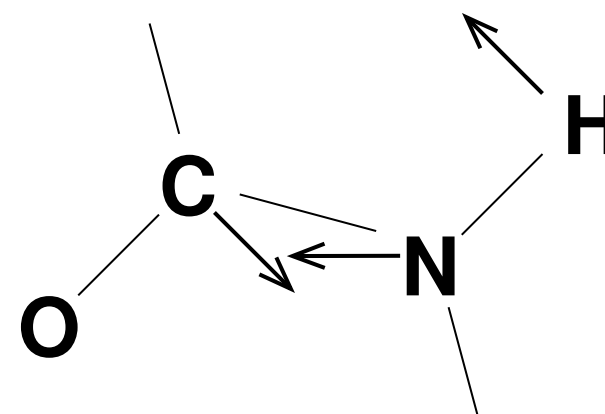




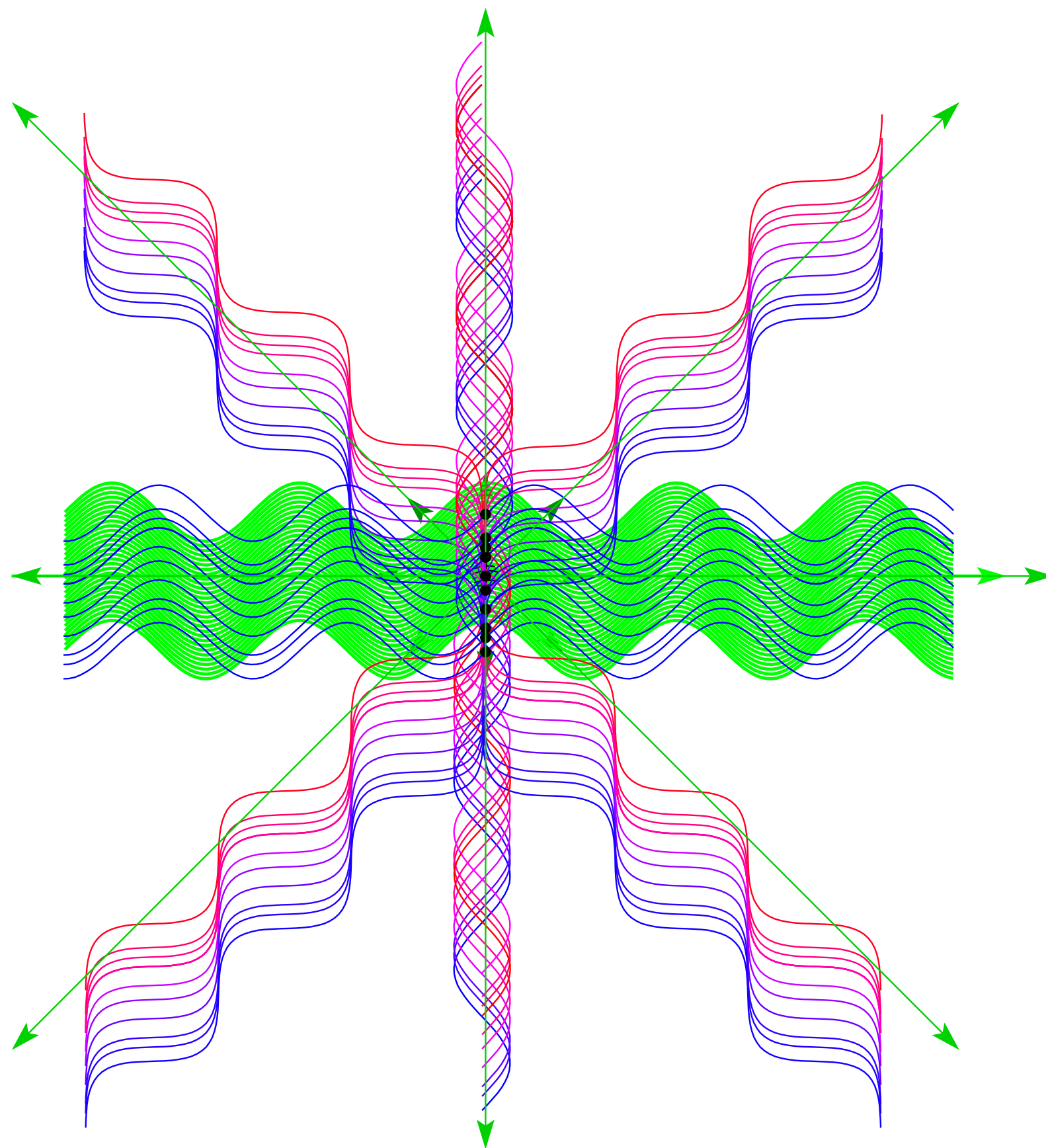
AMID I

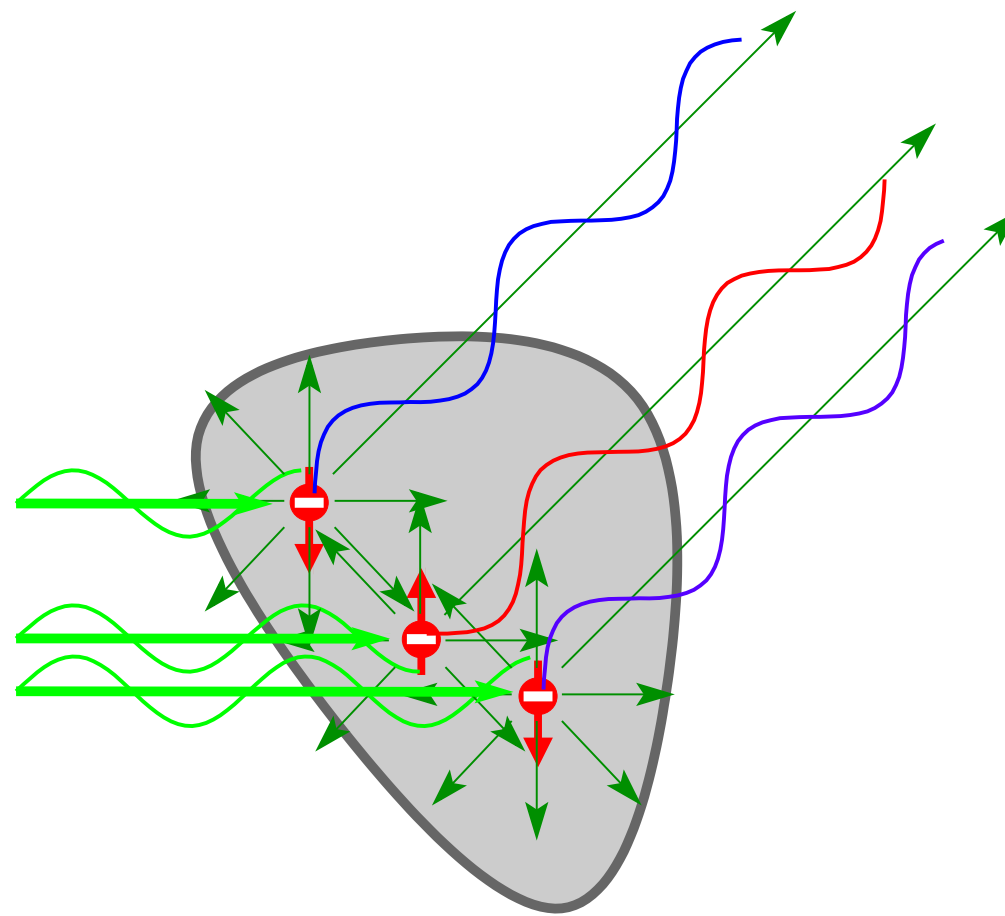


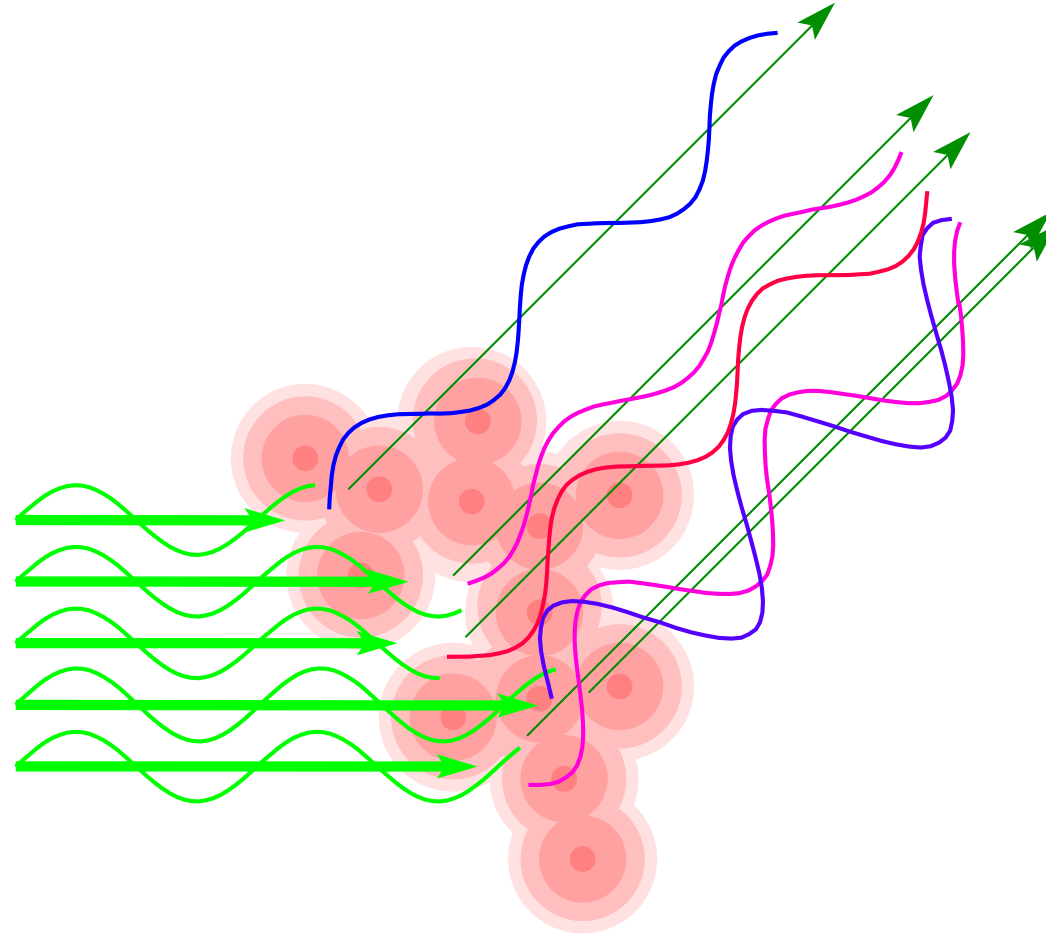
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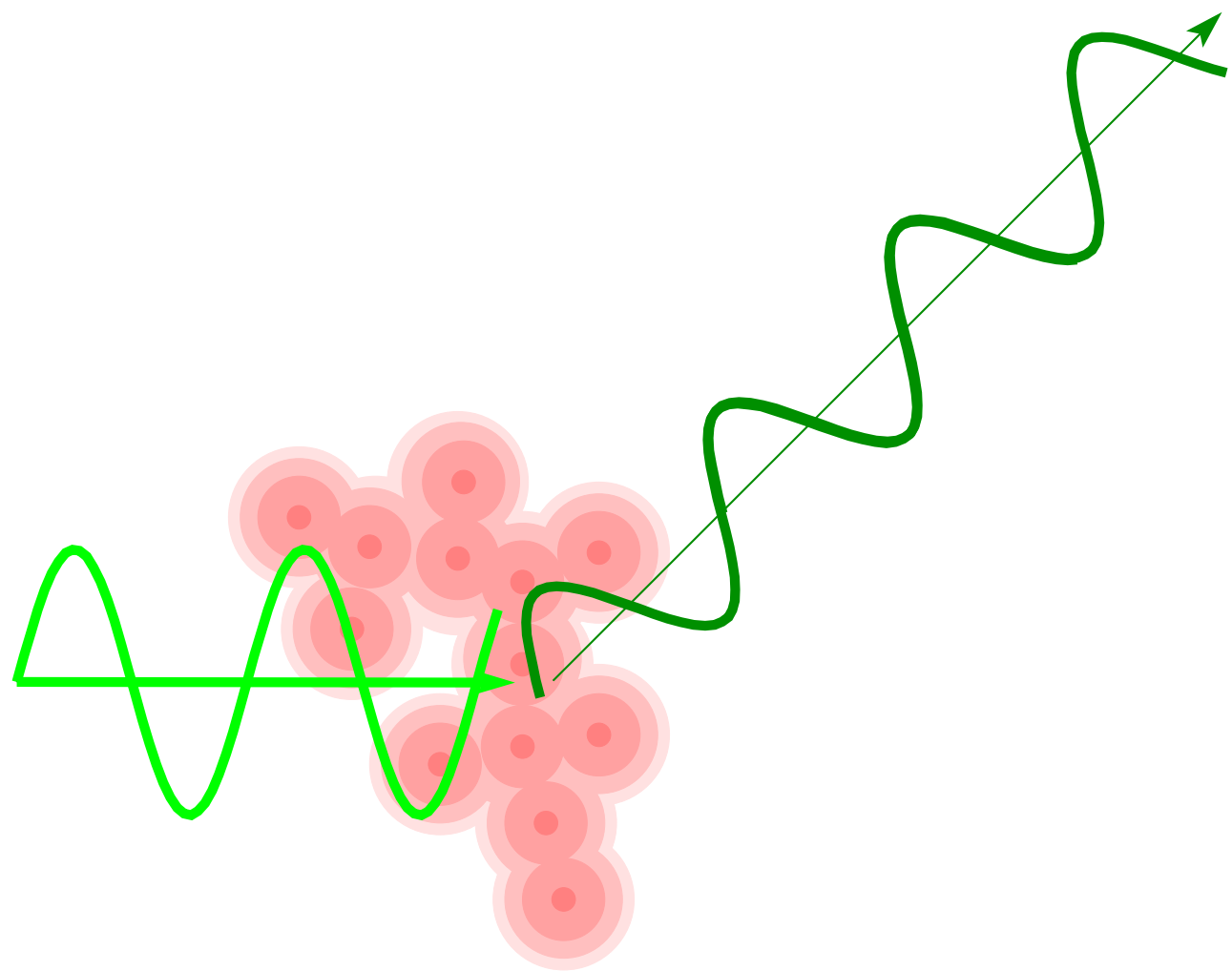


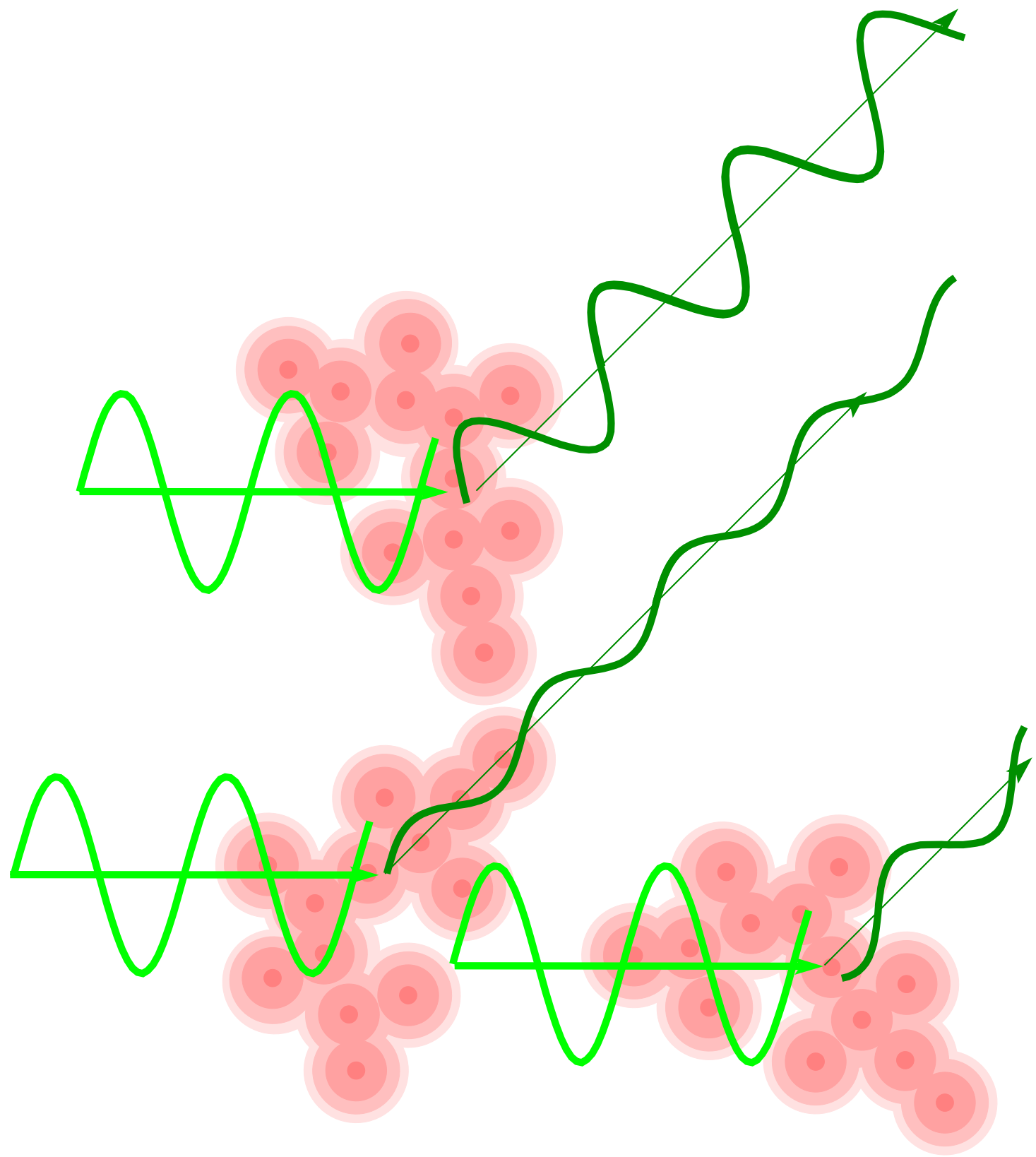
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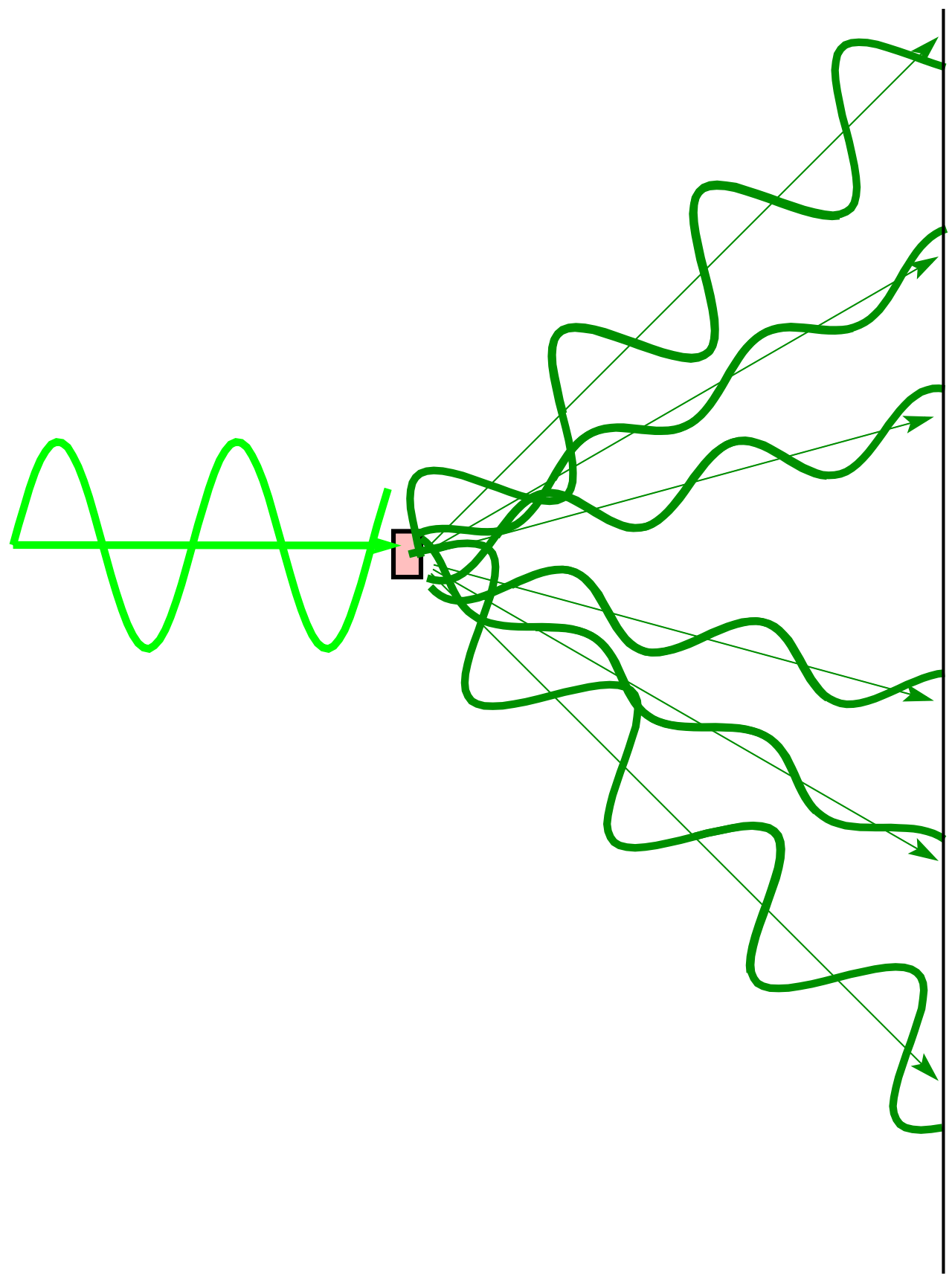


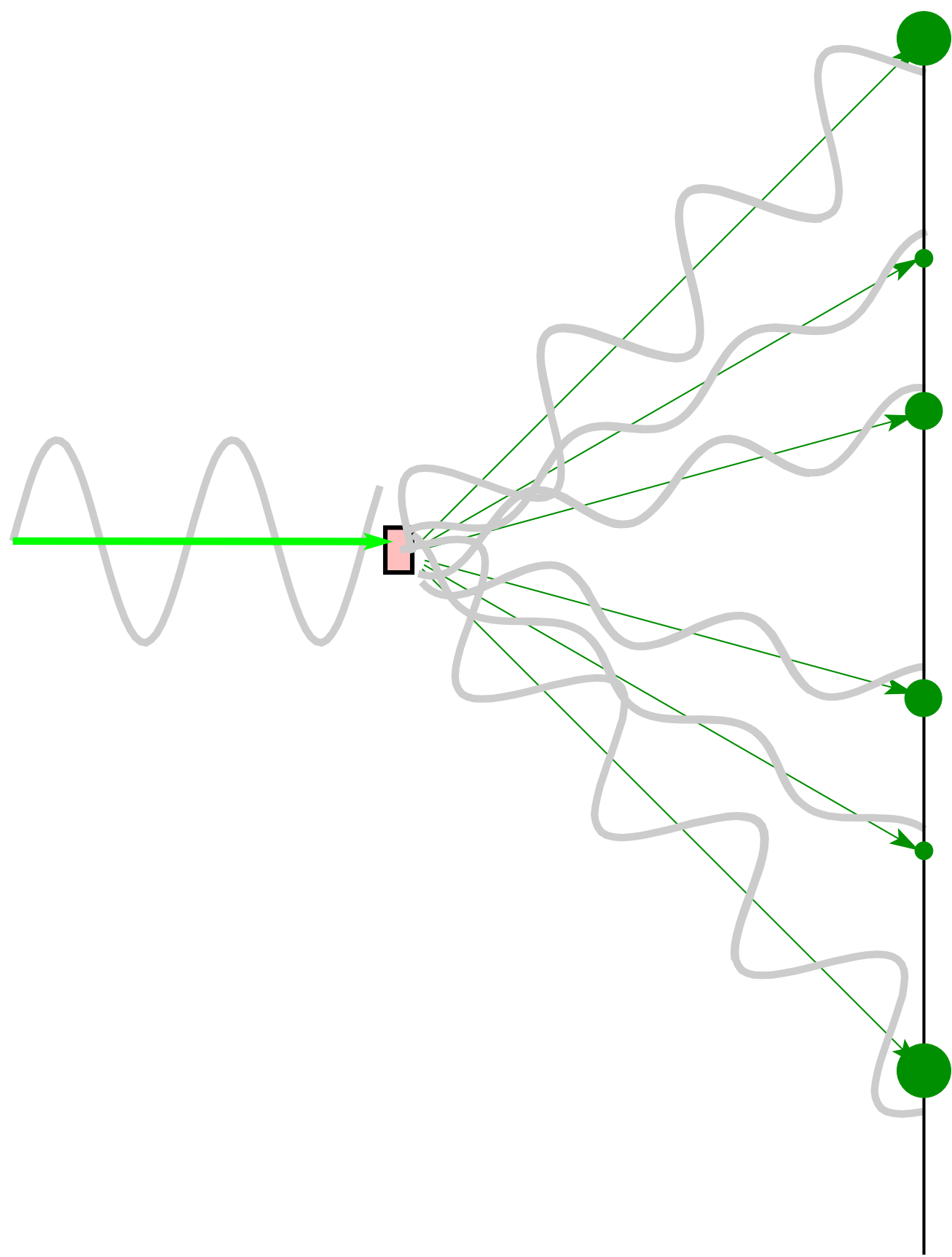




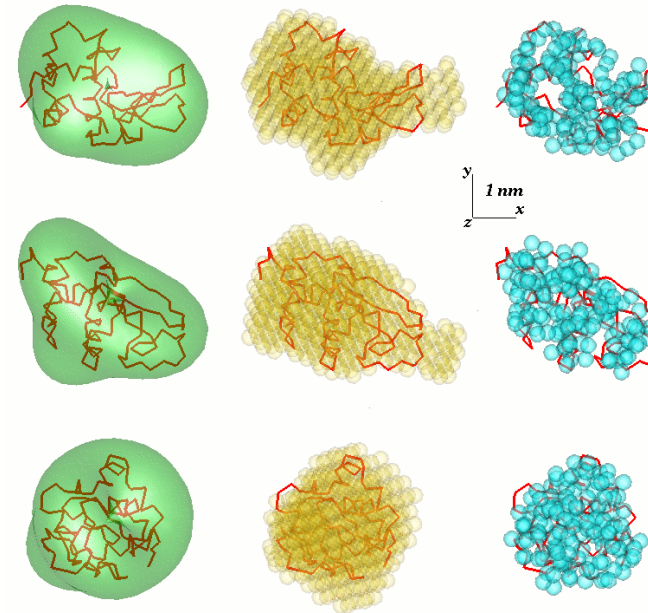
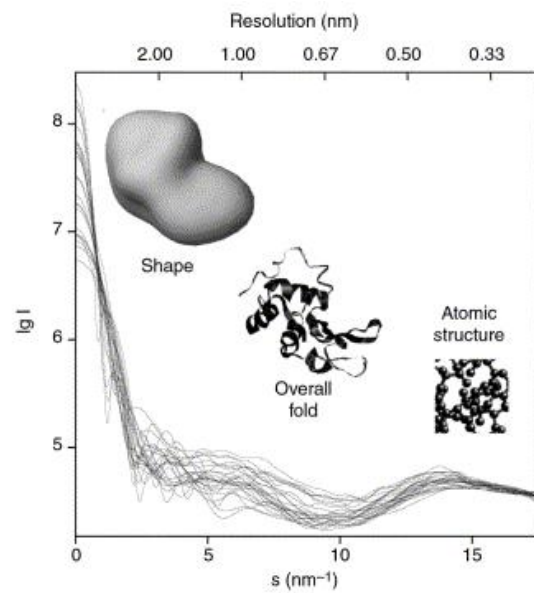








$$q = s = \frac{4\pi}{\lambda} \sin \theta$$



from Wikipedia

